

Kinetic Molecular Theory

50 Sample Test Questions

Multiple Choice

Identify the choice that best completes the statement or answers the question.

1. According to the kinetic theory, collisions between molecules in a gas _____.
A) are perfectly elastic
B) are inelastic
C) cause a loss of total kinetic energy
D) never occur
2. Which of the following statements is part of the kinetic theory?
A) The particles of a gas move independently of each other.
B) The particles in a gas move rapidly.
C) The particles in a gas are relatively far apart.
D) all of the above
3. The average speed of oxygen molecules in air is about _____.
A) 1700 km/h
B) 17,000 km/h
C) 0 km/h
D) 170 km/h
4. Particles in a gas are best described as _____.
A) spheres that are in fixed positions when trapped in a container
B) slow-moving, kinetic, hard spheres
C) hard spheres influenced by repulsive forces from other spheres
D) small, hard spheres with insignificant volumes
5. What is the SI unit of pressure?
A) pascal
B) newton
C) candela
D) mole
6. Standard conditions when working with gases are defined as _____.
A) 0°C and 101.3 kPa
B) 0°C and 1 kPa
C) 0 K and 1 kPa
D) 0 K and 101.3 kPa
7. How does the atmospheric pressure at altitudes below sea level compare with atmospheric pressure at sea level?
A) The atmospheric pressure below sea level is lower.
B) The atmospheric pressure below sea level is higher.
C) Differences in pressures cannot be determined.
D) The pressures are the same.
8. What causes gas pressure in a container such as a helium balloon?
A) atmospheric pressure acting on the outside walls of the container
B) the walls of the container
C) the vacuum maintained in the container
D) the simultaneous collisions of fast-moving particles in the container
9. The pressure of a gas in a container is 152 mm Hg. This is equivalent to _____.
A) 0.4 atm
B) 0.3 atm
C) 2 atm
D) 0.2 atm
10. The temperature at which the motion of particles theoretically ceases is _____.
A) 273°C
B) -273 K
C) 0°C
D) 0 K
11. What happens to the average kinetic energy of the particles in a sample of matter as the temperature of the sample is increased?
A) The average kinetic energy does not change.
B) The change in average kinetic energy cannot be determined.
C) The average kinetic energy decreases.
D) The average kinetic energy increases.

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The Princeton Review



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