

Kinetic-Molecular Theory

1. Each of these flasks are the same size and at the same temperature. Which flask contains the most molecules?



Flask # _____

How did you determine which flask?

2. Which of the following changes to a system WILL NOT result in an increase in pressure? Circle your choice.

1. increasing the volume of the container
2. decreasing the volume of the container
3. raising the temperature
4. adding more gas molecules

3. Each of these flasks contains the same number of gas particles. Which flask will have the lowest pressure?



Flask # _____

How did you determine which flask?

4. Each of these flasks contains the same number of gas particles. Which flask will have the lowest pressure?



Flask # _____

How did you determine which flask?

Kinetic-Molecular Theory of Gases

5. Gas particles are constantly in motion, moving in straight lines.

They are elastic collisions.

They have energy.



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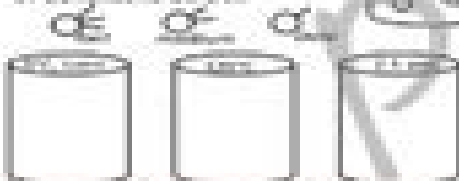
According to the Kinetic-Molecular Theory of Gases, above, that _____, _____, and _____.

Which five statements above do not apply to "real" gases?

1. _____

Which statements not apply to "real" gases?

Flask 1 contains 7 gas particles. Represent these same particles under the change of conditions below.



10. Under what conditions do "real" gases deviate the most from the ideal gas? Explain why.



Chemistry Corner

Kinetic Molecular Theory Of Gases Answer Key

Y Pai



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