F332 advance notice (June 2013) "predicted Q's"

- How does photolysis play a part in the destruction of N₂O in the stratosphere?
- Use a series of equations to show how NO acts in the catalytic destruction of ozone?
- What is meant by the term photolysis?
- Convert 100 parts per billion to a percentage, for the level of NO_x in urban areas.
- Why is the N₂O man into N₂O mor a true dynamic equilibrium?
- Describe how photochemical smog can be harmful to:
 - (a) the environment?
- (b) Humans ?
- The article mentions a reaction with HO radicals, what is a radical?
 Give examples of other radicals found in the stratosphere?
- Suggest why the concentration of N₂O in the atmosphere has increased since 1820 ?
- What is used as a catalyst in the upper atmosphere in photolytic fission leading to radicals being formed?
- Explain why N₂O is considered a greenhouse gas?
- Give the symbol equation for the reaction taking place in a catalytic converter to turn nitrogen oxides (NO_x) to oxygen and nitrogen?
 Is this oxidation or reduction?
- "Nitrous oxide is released into the atmosphere from the soil aquatic systems". Discuss?
- Why are the three bonds in nitrogen (N₂) significant?
- Explain using examples the mechanism of a radical chain reaction?
- What are the dangers of photochemical smog?
- Explain why lightning contributes to the formation of NO_x give one of the origins of NO_x?
- Draw a dot and cross bonding diagram for N₂O
- Loss of N₂O by wet deposition contributes to which atmospheric problem?
- Give the oxidation state of <u>N</u> in each of the following:
 - a) N₂O
- b) HNO₃
- e) NO
- d) (NH₄)₅SO₄
- Use a diagram to show the bonding in Ozone (O₁)
 Describe why the bonds aren't true double bonds?
- Write balanced symbol equations to represent the reactions in the use of a three-way catalyst?
- Describe how HNO₃ is removed from that atmosphere by wet deposition?

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