

F325 Acids, Bases and Buffers test ms

1. (a) completely dissociates/ionised (1) 2
proton donor (1)
- (b) NO_3^- (1) 1
- (c) (i) $\text{pH} = -\log[\text{H}^+] / -\log(0.015)$ (1) = 1.82 / 1.8 (1) (Not 2) 2
(ii) $[\text{H}^+] = 0.0075 \text{ mol dm}^{-3}$
 $\text{pH} = -\log(0.0075) = 2.12 / 2.1$ (1) 1
- [6]**
2. (a) (i) H^+ /proton donor (1) 1
(ii) partially dissociates/ionises (1) 1
- (b) $\text{C}_6\text{H}_5\text{OH}(\text{aq}) + \text{OH}^-(\text{aq}) \rightleftharpoons \text{C}_6\text{H}_5\text{O}^-(\text{aq}) + \text{H}_2\text{O}(\text{l})$
acid 1 base 2 (1) base 1 acid 2 (1)
1 mark for each acid-base pair
- (c) (i) $K_a = \frac{[\text{C}_6\text{H}_5\text{O}^-][\text{H}^+]}{[\text{C}_6\text{H}_5\text{OH}]}$ (1) 1
(ii) concentration = $38/94$ (1) = 0.40 mol dm^{-3} (1)
(first mark for M_r of phenol – incorrect answer here will give ecf for remainder of question)
- $$1.3 \times 10^{-10} \approx \frac{[\text{H}^+(\text{aq})]^2}{0.40} \quad (1)$$
- (‘=’ sign is acceptable)
- $[\text{H}^+] = \sqrt{((1.3 \times 10^{-10}) \times (0.40))} = 7.2 \times 10^{-6} \text{ mol dm}^{-3}$ (1)
 $\text{pH} = -\log[\text{H}^+] = -\log 7.2 \times 10^{-6} = 5.14$ (1) 5
- 3 marks: $[\text{H}^+]$ (1), pH expression (1), calc of pH from $[\text{H}^+]$ (1)
- Common errors:**
Without square root, answer = 10.28 (1)(1)(0)
Use of 38 as molar concentration does not score 1st 2 marks.
This gives an answer of 4.15 for 3 marks (1)(1)(1)
- (d) $[\text{H}^+(\text{aq})] = 1.99 \times 10^{-6} \text{ mol dm}^{-3}$ (1)
 $[\text{C}_6\text{H}_5\text{O}^-(\text{aq})] = K_a [\text{C}_6\text{H}_5\text{OH}(\text{aq})] / [\text{H}^+(\text{aq})]$ (1)
 $[\text{C}_6\text{H}_5\text{O}^-(\text{aq})] = 0.013 \text{ mol dm}^{-3}$ (1)
- Calculation should use half the original concentration of phenol to find the concentration of sodium phenoxide in the buffer. This should then be doubled back up again.*
Do not penalise an approach that uses the original concentration of phenol in the expression above.

[13]

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L Cohen



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