Limiting Reagent & Percent Yield Practice Worksheet

- When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.
 - a. Write the balanced equation for the reaction given above:

 $CuCl_2 + NaNO_3 \rightarrow Cu(NO_3)_2 + NaCl$ $CuCl_2 + 2 NaNO_3 \rightarrow Cu(NO_3)_2 + 2 NaCl$

b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

15 g CuCl₂ x 1 mol CuCl₂ x 2 mol NaCl x 58.443 g NaCl = 13 g NaCl 134.362g CuCl₂ 1 mol CuCl₂ 1 mol NaCl

20 g NaNO₃ x <u>1 mol NaNO₃</u> x <u>2 mol NaCl</u> x <u>58.443 g NaCl</u> = 13.8 g NaCl 84.997g NaNO₃ 2 mol NaNO₃ 1 mol NaCl

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (11) chloride
- d. How many grams of copper (II) nitrate is formed?

15g CuCl₂ x 1 mol CuCl₂ x 1 mol Cu(NO₃)₂ x 187.467 g Cu(NO₃)₂ = 20.9 g Cu(NO₃)₂ = 134.452g CuCl₂ 1 mol CuCl₂ 1 mol Cu(NO₃)₂

e. How much of the excess reagent is left over in this reaction?

15g CuCl₂ x 1 mol CuCl₂ x 2 mol NaNO₃ x 84,997g NaNO₃ = 19 g NaNO₃ 134,452g CuCl₂ 1 mol CuCl₂ 1 mol NaNO₃

20 g NaNO₅ - 19 g NaNO₅ = 1 g NaNO₅

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction? Percent yield = actual yield / theoretical yield × 100 percent

11.3/13.0 x 100% = 86.9%

- When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.
 - a. Balance the following equation:

 $Pb(NO_3)_{2 \text{ (sq)}} + NaI_{\text{ (sq)}} \rightarrow PbI_{2 \text{ (s)}} + NaNO_{3 \text{ (sq)}}$ $Pb(NO_3)_{2 \text{ (sq)}} + 2 NaI_{\text{ (sq)}} \rightarrow PbI_{2 \text{ (s)}} + 2 NaNO_{3 \text{ (sq)}}$

b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

25.0 g Pb(NO₃)₂ x 1 mol Pb(NO₃)₂ x 2 mol NaNO₃ x 84.997g NaNO₃ = 12.8 g NaNO₃ 331.214 g Pb(NO₃)₂ 1 mol Pb(NO₃)₂ 1 mol NaNO₃

15.0 g Nal x 1 mol Nal x 2 mol NaNO₃ x 84.997g NaNO₃ = 8.51 g NaNO₃ 149.89g Nal 2 mol Nal 1 mol NaNO₃

8.51 g NaNO3 can be made

Limiting Reagent And Percent Yield 9 3 Answers

Michael Hoelscher

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