

Limiting Reactant and Percent Yield Worksheet

Key

1. Calcium hydroxide, used to neutralize acid spills, reacts with hydrochloric acid according to the following unbalanced equation:



- a. If you have spilled 6.3 mol of HCl and put 2.8 mol of Ca(OH)_2 on it, which substance is the limiting reactant?

$$\frac{6.3 \text{ mol HCl}}{2 \text{ mol HCl}} \bigg| \frac{1 \text{ mol CaCl}_2}{2 \text{ mol HCl}} = 3.2 \text{ mol CaCl}_2$$

$$\frac{2.8 \text{ mol Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \bigg| \frac{1 \text{ mol CaCl}_2}{1 \text{ mol Ca(OH)}_2} = 2.8 \text{ mol CaCl}_2$$

* Ca(OH)_2 is limiting

- b. How many moles of the excess reactant remain?

$$\frac{2.8 \text{ mol Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \bigg| \frac{2 \text{ mol HCl}}{1 \text{ mol Ca(OH)}_2} = 5.6 \text{ mol HCl used}$$

$$6.3 \text{ mol HCl} - 5.6 \text{ mol HCl} = \boxed{0.7 \text{ mol HCl remain}}$$

2. Aluminum oxidizes according to the following unbalanced equation: $\underline{4}\text{Al} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{Al}_2\text{O}_3$

- a. Powdered Al (0.048 mol) is placed into a container containing 0.030 mol O_2 . What is limiting reactant?

$$\frac{0.048 \text{ mol Al}}{4 \text{ mol Al}} \bigg| \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} = 0.024 \text{ mol Al}_2\text{O}_3$$

$$\frac{0.030 \text{ mol O}_2}{3 \text{ mol O}_2} \bigg| \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol O}_2} = 0.020 \text{ mol Al}_2\text{O}_3$$

* O_2 is limiting

- b. How many moles of the excess reactant remain?

$$\frac{0.030 \text{ mol O}_2}{3 \text{ mol O}_2} \bigg| \frac{4 \text{ mol Al}}{3 \text{ mol O}_2} = 0.040 \text{ mol Al used}$$

$$0.048 \text{ mol Al} - 0.040 \text{ mol Al} = \boxed{0.008 \text{ mol Al remain}}$$

3. Heating zinc sulfide in the presence of oxygen yields the following: $\underline{2}\text{ZnS} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{ZnO} + \underline{2}\text{SO}_2$

- a. If 1.72 mol of ZnS is heated in the presence of ^{45.6}~~32.0~~ g of O_2 , which is the limiting reactant? (Balance the equation first.)

$$\frac{1.72 \text{ mol ZnS}}{1 \text{ mol ZnS}} \bigg| \frac{2 \text{ mol SO}_2}{1 \text{ mol ZnS}} = 3.44 \text{ mol SO}_2$$

$$\frac{45.6 \text{ g O}_2}{32.0 \text{ g O}_2} \bigg| \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \bigg| \frac{2 \text{ mol SO}_2}{3 \text{ mol O}_2} = 0.950 \text{ mol SO}_2$$

* O_2 is limiting

- b. How many grams of the excess reactant remain?

$$\frac{45.6 \text{ g O}_2}{32.0 \text{ g O}_2} \bigg| \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \bigg| \frac{2 \text{ mol ZnS}}{3 \text{ mol O}_2} \bigg| \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 92.6 \text{ g used}$$

$$\frac{1.72 \text{ mol ZnS}}{1 \text{ mol ZnS}} \bigg| \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 168 \text{ g ZnS (started with)}$$

$$\begin{array}{r} 168 \text{ g ZnS} \\ - 92.6 \text{ g ZnS} \\ \hline 75 \text{ g ZnS} \\ \text{remain} \end{array}$$

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R Bogdan



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