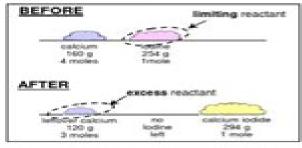
Name _____

When an automobile production plant runs out of tires no more cars can be produced even if there are still plenty of engines, bodies, seats, and other parts. When one of the reactants in a chemical reaction runs out the reaction stops as well. The substance that limits production in a chemical reaction is known as the **limiting** reactant. The reactant that is left over when the reaction stops is called the excess reactant.

1 Ca + 1 I₂ → 1 CaI₂

Consider the chemical reaction between calcium and iodine. When 160 g (4 moles) of calcium reacts with 254 g (1 mole) of iodine the reaction makes 294 g (1 mole) of calcium iodide. Although there is enough calcium to make more calcium iodide, the iodine runs out first. Since all of the iodine gets used up it is called the limiting reactant. According to the balanced equation, if one mole of iodine reacts, one mole of calcium will react. This means that there are still 3 moles of calcium left. Because calcium is left over it is called the excess reactant.



Example

What mass of iron (II) sulfide will be produced if 9.68 g of iron reacts with 6.28 g of sulfur? __Fe + __S → __FeS

- balance the equation <u>1</u> Fe + <u>1</u> S → <u>1</u> FeS
- perform a calculation for each reactant: $p: \frac{9.68 \text{ g/Gc}}{1} \times \frac{1 \text{ m/h/fc}}{55.85 \text{ g/Gc}} \times \frac{1 \text{ m/h/fc}}{1 \text{ m/h/fc}} \times \frac{87.91 \text{ g/Fc}}{1 \text{ m/h/fc}} = 15.24 \text{ g/Fc}$
 - S: 6.28 g/s 1 miles 1 mol Fes | 87.91 g FeS | 17.22 g FeS | 1 mol Fes | 1 mol Fes |
- whichever reactant makes less product is
 the limiting reactant. There is enough sulfur to make 17.22 g
 feS, but only enough iron to make 15.24 g of FeS.

Find the amount of product formed in each of the following reactions.

- What mass of H₂O will be produced if 9.5 g of H₂ reacts with 1.2 g of O₂?
 ___ H₂ + ___ O₂ → ___ H₂O
- If 1.85 g of Mg(OH)₂ reacts with 3.71 g of HCl, how much MgCl₂ is produced? What is the limiting reactant? __Mg(OH)₂ + __HCl → __MgCl₂ + __H₂O
- What mass of AgCl is produced when 53.42 g of AgNO₃ reacts with 14.19 g of NaCl? __AgNO₃ + __NaCl → __AgCl + __NaNO₃
- If 14.7 g of calcium is placed in 11.5 g of water, what mass of hydrogen gas is produced? __Ca + __H₂O → __Ca(OH)₂ + __H₂
- What mass of potassium hydroxide is formed when 8.2 g of potassium oxide is added to 1.3 g of water? ___ K₂O + __ H₂O → __ KOH

- What mass of aluminum chloride could be made from 8.1 g of aluminum and 4.2 L of chlorine at STP? __ Al + __ Cl₂ → __ AlCl₃
- If 5.26 L of nitrogen monoxide and 7.64 L of oxygen are combined, what mass of nitrogen dioxide is formed? Assume conditions are STP.
 NO + _ O₂ → _ NO₂
- If 18.1 g of silicon tetrachloride reacts with 8.4 L of hydrogen at STP, what mass of silicon is formed? __SiCl₄ + __H₂ → __Si + __HCl
- What mass of Cu(OH)₂ is produced in the reaction of 7.6 g Cu(NO₃)₂ with 6.2 g of KOH? _ Cu(NO₃)₂ + _ KOH → _ Cu(OH)₂ + _ KNO₃
- What mass of carbon dioxide is formed when 64 kg of ethylene is burned in 142 kg of oxygen?
 _ C₂H₄ + _ O₂ → _ CO₂ + _ H₂O

Limiting Reactions Chem Worksheet 12 3 Answers Key

Arturo Cuomo

Limiting Reactions Chem Worksheet 12 3 Answers Key:

Unveiling the Power of Verbal Artistry: An Psychological Sojourn through **Limiting Reactions Chem Worksheet 12 3 Answers Key**

In a world inundated with monitors and the cacophony of immediate conversation, the profound power and emotional resonance of verbal artistry usually disappear in to obscurity, eclipsed by the continuous onslaught of sound and distractions. Yet, nestled within the musical pages of **Limiting Reactions Chem Worksheet 12 3 Answers Key**, a captivating perform of literary splendor that pulses with organic feelings, lies an remarkable trip waiting to be embarked upon. Written by a virtuoso wordsmith, that mesmerizing opus instructions viewers on an emotional odyssey, softly revealing the latent possible and profound impact embedded within the complex web of language. Within the heart-wrenching expanse of this evocative examination, we will embark upon an introspective exploration of the book is key subjects, dissect their interesting writing fashion, and immerse ourselves in the indelible effect it leaves upon the depths of readers souls.

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