

Activity B:**Methane reaction**Get the Gizmo ready:

- Select $C + H_2$ becomes CH_4 .



Goal: Predict the results of a chemical reaction.

1. Count: CH_4 is the chemical formula for methane. In order to produce a single molecule of methane, how many carbon (C) atoms are needed? Hydrogen atoms?

2. Predict: Set the number of C atoms to 14 and the number of H_2 molecules to 10.

A. How many total hydrogen atoms are there?

B. How many methane molecules could form from these reactants?

C. After the atoms react, which reactant will be left over?

D. Which reactant will be the limiting reactant?

E. Click **Play** and wait until **Reaction complete** is shown. What happened?

Each hydrogen molecule formed a connection with carbon (4 hydrogen molecules per carbon). There are extra carbon atoms that did not link because there are additional carbon atoms.

3. Test: Click **Reset**. Set the number of C atoms to seven.

A. How many hydrogen molecules (H_2) will be needed for there to be no excess reactants? Use the slider to set the number of H_2 molecules to this value.

B. How many methane molecules will be formed?

C. Click **Play**. What happened?

Since all of the hydrogen molecules linked to the carbon, there were no extra reactants.

4. Solve: For each **chemical equation** below, write the number of product molecules that will form from the reaction. Then, circle the limiting reactant. (Note: The **coefficients** in front of the reactants indicate the number of reactant molecules or atoms present.)



Limiting Reactants Gizmo Answer Key

Mark Steyvers

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Limiting Reactants Gizmo Answer Key

1. Understanding the eBook Limiting Reactants Gizmo Answer Key
 - The Rise of Digital Reading Limiting Reactants Gizmo Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactants Gizmo Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactants Gizmo Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactants Gizmo Answer Key
 - Personalized Recommendations
 - Limiting Reactants Gizmo Answer Key User Reviews and Ratings
 - Limiting Reactants Gizmo Answer Key and Bestseller Lists
5. Accessing Limiting Reactants Gizmo Answer Key Free and Paid eBooks
 - Limiting Reactants Gizmo Answer Key Public Domain eBooks
 - Limiting Reactants Gizmo Answer Key eBook Subscription Services
 - Limiting Reactants Gizmo Answer Key Budget-Friendly Options

6. Navigating Limiting Reactants Gizmo Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reactants Gizmo Answer Key Compatibility with Devices
 - Limiting Reactants Gizmo Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactants Gizmo Answer Key
 - Highlighting and Note-Taking Limiting Reactants Gizmo Answer Key
 - Interactive Elements Limiting Reactants Gizmo Answer Key
8. Staying Engaged with Limiting Reactants Gizmo Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactants Gizmo Answer Key
9. Balancing eBooks and Physical Books Limiting Reactants Gizmo Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactants Gizmo Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactants Gizmo Answer Key
 - Setting Reading Goals Limiting Reactants Gizmo Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactants Gizmo Answer Key
 - Fact-Checking eBook Content of Limiting Reactants Gizmo Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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