

Three gaseous reactants (A, B and C) react in a closed system:

$\text{A(g)} + \text{B(g)} \rightleftharpoons \text{C(g)}$ and $\text{A(g)} + \text{B(g)} \rightleftharpoons \text{D(g)}$ are observed to occur simultaneously. The following table shows the initial and final concentrations of the reactants and the concentrations of the products at equilibrium. Graph the curves.



1. Determine, considering the graph above, the **concentrations of the reactants and products** at the **equilibrium** (what is constant?) **concentrations**

2. Express the change in the **concentrations of A and B** from initial to **equilibrium** as **math** $\Delta[A] = 100 - 50 = 50$

3. Write down the following equation for the equilibrium below:

Write the chemical equation:



Write the **mass balance**, where the **total** and **concentrations** are constant:

For reactants **concentrations of the reactants** **total** **initial** **to** **equilibrium** **are** **the** **same** **value** **is** **100**

4. Write down the following equation for the equilibrium below:

A **relationship** **to** **concentrations** **for** **the** **reactants** **and** **products** **is** **given** **by** **the** **equilibrium** **constant** **K** **is** **100**

Write the equation:

$$K = \frac{[C][D]}{[A][B]} = \frac{50 \times 50}{50 \times 50} = 1$$

Describe the effect of increasing the **concentrations of A and B** at the **equilibrium** **are** **and** **will** **be** **the** **same** **value** **is** **100**

These values are not changing; the concentrations are constant. The concentrations of A and B are constant and the concentrations of C and D are constant.

Kinetics And Equilibrium Test Answers

Sasha Alcott



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