

CHEMISTRY

CHEAT SHEET

CHEMICAL KINETICS

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DEFINITIONS: The study of the rates of chemical reactions is called chemical kinetics. It is concerned with the rates of chemical reactions and the factors that affect them.

Rate of Reaction: The rate of a chemical reaction is the change in concentration of a reactant or product per unit time.

Factors Affecting Rate: Temperature, concentration, surface area, catalyst, and physical state.

Rate Law: An equation that relates the rate of a reaction to the concentrations of the reactants.

Order of Reaction: The sum of the powers of the concentration terms in the rate law.

Half-life: The time required for the concentration of a reactant to decrease to half its initial value.

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Reaction Orders			
Order	Description	Rate Law	Half-life
Zero-order	Rate is independent of the concentration of reactants	$\text{Rate} = k$	$t_{1/2} = \frac{[A]_0}{k}$
First-order	Rate is directly proportional to the concentration of one reactant	$\text{Rate} = k[A]$	$t_{1/2} = \frac{0.693}{k}$
Second-order	Rate is proportional to the square of the concentration of one reactant or to the product of the concentrations of two reactants	$\text{Rate} = k[A]^2$ or $\text{Rate} = k[A][B]$	$t_{1/2} = \frac{1}{k[A]_0}$

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Kinetics Study Guide

Thomas Griffiths



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