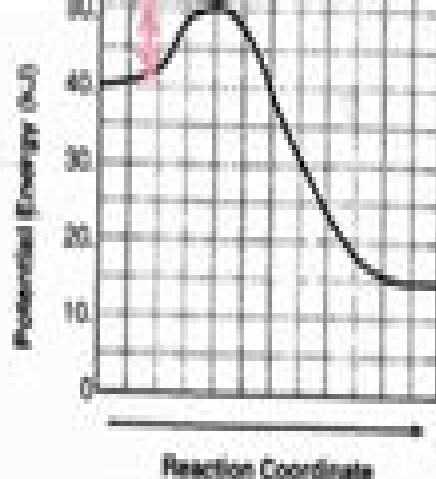




Which statement correctly describes the energy changes that occur in the forward reaction?

- ☐ A) The activation energy is 10. kJ and the reaction is endothermic.
- ☒ B) The activation energy is 10. kJ and the reaction is exothermic.
- ☐ C) The activation energy is 50. kJ and the reaction is endothermic.
- ☐ D) The activation energy is 50. kJ and the reaction is exothermic.

9. Given the reaction at equilibrium:



The equilibrium point will shift to the right if the pressure is

- ☐ A) increased and the temperature is increased
- ☒ B) increased and the temperature is decreased

- formation of 1 mole of $H_2O(l)$)
- ☒ A) -285.8 kJ
 - ☐ B) $+285.8 \text{ kJ}$
 - ☐ C) -571.6 kJ
 - ☐ D) $+571.6 \text{ kJ}$

11. As the number of effective collisions between reacting particles increases, the rate of reaction

- ☐ A) decreases
- ☒ B) increases
- ☐ C) remains the same

12. Which statement must be true when solution equilibrium occurs?

- ☐ A) The solution is at STP.
- ☐ B) The solution is supersaturated.
- ☒ C) The concentration of the solution remains constant.
- ☐ D) The masses of the dissolved solute and the undissolved solute are equal.

13. Given the reaction at equilibrium:



The addition of a catalyst will

- ☐ A) shift the equilibrium to the right
- ☐ B) shift the equilibrium to the left
- ☒ C) increase the rate of forward and reverse reactions equally
- ☐ D) have no effect on the forward or reverse reactions

14. Which equation shows an increase in entropy?

- ☐ A) $CO_2(g) \rightarrow CO_2(s)$
- ☒ B) $CO_2(l) \rightarrow CO_2(g)$

Kinetics And Equilibrium Practice Test

The Princeton Review



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