

## PHYSICAL: Kinetic Molecular Theory

### Learning Objectives

- At standard atmospheric conditions, all gases have approximately the same kinetic energy, but are made of different molecules and are made up of different numbers of particles. (The kinetic molecular theory is based on the assumption that all particles have the same average kinetic energy.)
- At standard atmospheric conditions, gases have the same volume per molecule.

### Model 1: Representation of Gases in Different Phases

The following figure illustrates the representation of particles in three phases: solid, liquid, and gas.

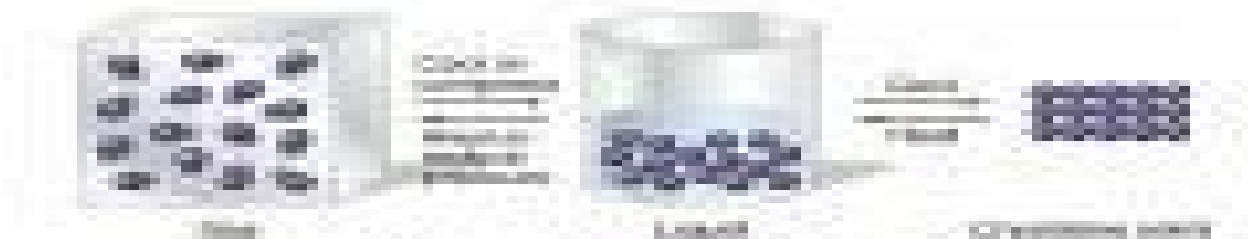


Figure 1: Representation of particles in different phases of matter

Model 1 represents the particles in different phases of matter. The particles in the solid phase are closely packed, while the particles in the liquid phase are more disordered. The particles in the gaseous phase are widely spaced.

### Key Concepts

1. Model 1 shows the representation of particles in different phases of matter. The particles in the solid phase are closely packed, while the particles in the liquid phase are more disordered. The particles in the gaseous phase are widely spaced.

Table 1: Representation of particles in different phases

|                                 | Solid                                                                                   | Liquid                                                                                  | Gaseous                                                                             |
|---------------------------------|-----------------------------------------------------------------------------------------|-----------------------------------------------------------------------------------------|-------------------------------------------------------------------------------------|
| How are the particles arranged? | The particles are closely packed together. The arrangement is regular.                  | The particles are closely packed together. The arrangement is disordered.               | The particles are widely spaced. The arrangement is disordered.                     |
| How are the particles moving?   | The particles are vibrating in place. They are not moving from one position to another. | The particles are moving from one position to another. They are not vibrating in place. | The particles are moving from one position to another. They are vibrating in place. |

# Kinetic Molecular Theory Pogil Instructors Guide

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Haight**



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Teachers.Net Lesson Exchange: Kinetic Molecular Theory with Popcorn , Cec Knight offers a lesson for middle and high school students that uses popcorn to demonstrate the kinetic molecular theory Knight includes a list of the materials required the time needed and the procedures Teachers Net provides the lesson as part of the Teachers Net Lesson Exchange online resource

Instructor's Manual for General Chemistry, Experiment and Theory Richard H. Eastman,1970      **Kinetic Molecular Theory** Charles Proteus Steinmetz,American Electrochemical Society. Meeting,1908\*      **Instructor's Guide for Experiments in General Chemistry** ,1996      **Instructor's Guide for Experiments in General Chemistry** Steven Murov,1996

*Instructor's Guide for Dickerson/Gray/Haight Chemical Principles* Ben B. Chastain,Richard Earl

Dickerson,Harry B. Gray,Gilbert Pierce Haight,1979      *Communities of Molecules* Howard DeVoe,Robert Hearle,1978

Instructor's Solutions Manual to Accompany Chemistry Alan J. Pribula,Silberberg, Frank R. Miliro,Thomas H.

Berg,1999-11-01      *Instructor's Guide for Chemical Principles* Richard Earl Dickerson,1984      Chemistry Steven S.

Zumdahl,Susan A. Zumdahl,Donald J. DeCoste,2007

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