

(i) Ionic bond (ii) Covalent bond (iii) Co-ordinate bond

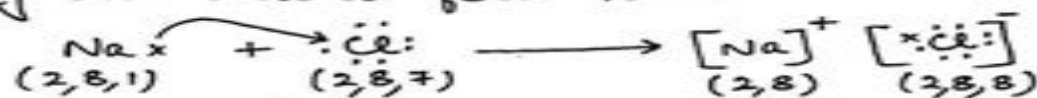
In addition to these 3 main types of chemical bonds, the atoms in a metal are linked by metallic bonds. While hydrogen bond and Vander Waals forces arise because of intermolecular forces of attraction in the covalent molecules.

* IONIC BOND OR ELECTROVALENT BOND :-

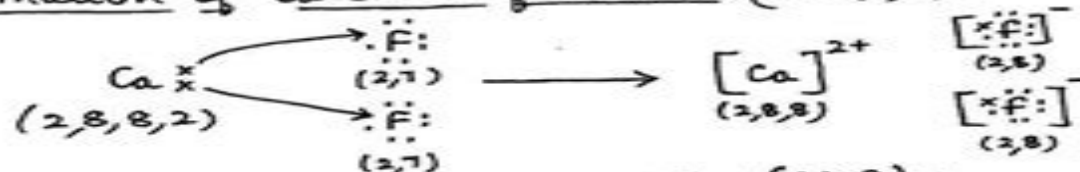
The bond formed by complete transference of e^- from one atom to another so that both of them acquire nearest noble gas configuration is called ionic bond. It is formed by atoms having large electronegativity difference. The atom which loses e^- acquires positive charge and which gains e^- acquires negative charge. The +vely and -vely charged ions are held together by the strong electrostatic force of attraction.

Examples :-

1. Formation of NaCl :- Sodium atom ($Z=11$) has one valence e^- (2, 8, 1). Highly Chlorine atom ($Z=17$) has 7 valence e^- (2, 8, 7). Sodium atom loses its only valence e^- and changes to Na^+ ion (cation). This e^- is taken up by Chlorine atom which changes to Cl^- ion (anion). But these ions are mutually attracted to form NaCl as shown below :-



(b) Formation of Calcium fluoride (CaF_2) :-



(c) Formation of magnesium oxide (MgO) :-



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