

Definitions

Atom -

Particles that make up all substances.

Element -

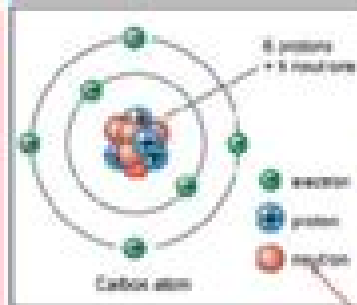
A substance made up of only one kind of atom.

Compound -

A substance made of different types of atoms joined together.

Mixture -

A substance consisting of atoms that can be easily separated as they are not joined together.



Periodic Table of the Elements

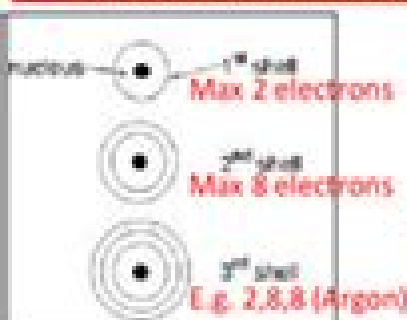
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K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra	Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr	

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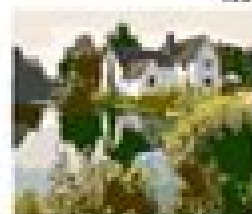
Group 0 or 8 are **UNREACTIVE**
They have a full outer shell of electrons

Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
E.g. Lithium 2,1 is a group 1 element.



Protons and neutrons both have a mass of 1



Used for building

LIMESTONE

Calcium Carbonate **CaCO₃**

HEAT

(thermal decomposition)

CaO

Calcium Oxide

Limewater

The test for CO₂

Ca(OH)₂

Or Calcium Hydroxide

comes from adding water to CaO

Carbonates
Decompose when heated
= Oxide + CO₂
React with acid
= Salt + water + CO₂

Balanced Equations



C=1

H=4

O=4

C=1

H=4

O=4

A balanced equation has the same number of atoms of each element on both sides

Using Molecules: In this equation, the large number in front of the chemical symbol tells you the molecules needed.

e.g. 1 CH₄ molecule reacts with 2 O₂ molecules

Metals - Very useful
e.g. Copper wires conduct well
Most need to be **ALLOYED** to make them harder

Metals which are un-reactive are found in their **NATIVE STATE** e.g. **GOLD**
More reactive metals are found as **METALS ORES** and need to be **EXTRACTED**

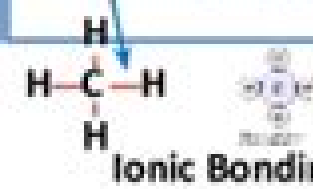


TRANSITION METALS



Covalent Bonding

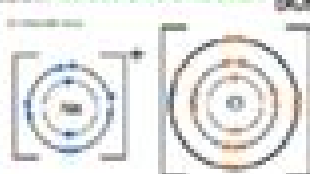
Non-metals sharing electrons



Ionic Bonding



IONS
Charged particles



C1 CHEMISTRY

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