

5 (d) Explain why the current in the external circuit of this cell falls to zero after the cell has operated for some time.

.....

.....

(1 mark)

5 (e) The simplified electrode reactions in a rechargeable lithium cell are

Electrode **A** $\text{Li}^+ + \text{MnO}_2 + \text{e}^- \longrightarrow \text{LiMnO}_2$ $E = -0.15 \text{ V}$

Electrode **B** $\text{Li}^+ + \text{e}^- \longrightarrow \text{Li}$

Electrode **B** is the negative electrode.

5 (e) (i) The e.m.f. of this cell is 2.90 V.

Use this information to calculate a value for the electrode potential of electrode **B**.

.....

.....

(1 mark)

5 (e) (ii) Write an equation for the overall reaction that occurs when this lithium cell is being recharged.

.....

.....

.....

(2 marks)

5 (e) (iii) Suggest why the recharging of a lithium cell may lead to release of carbon dioxide into the atmosphere.

.....

.....

(1 mark)

Gcse Chemistry Exam Questions And Answers

Mark Steyvers



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Table of Contents Gcse Chemistry Exam Questions And Answers

1. Understanding the eBook Gcse Chemistry Exam Questions And Answers
 - The Rise of Digital Reading Gcse Chemistry Exam Questions And Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gcse Chemistry Exam Questions And Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gcse Chemistry Exam Questions And Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gcse Chemistry Exam Questions And Answers
 - Personalized Recommendations
 - Gcse Chemistry Exam Questions And Answers User Reviews and Ratings
 - Gcse Chemistry Exam Questions And Answers and Bestseller Lists
5. Accessing Gcse Chemistry Exam Questions And Answers Free and Paid eBooks
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 - Gcse Chemistry Exam Questions And Answers eBook Subscription Services
 - Gcse Chemistry Exam Questions And Answers Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Gcse Chemistry Exam Questions And Answers Compatibility with Devices
 - Gcse Chemistry Exam Questions And Answers Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Gcse Chemistry Exam Questions And Answers
 - Highlighting and Note-Taking Gcse Chemistry Exam Questions And Answers
 - Interactive Elements Gcse Chemistry Exam Questions And Answers
8. Staying Engaged with Gcse Chemistry Exam Questions And Answers

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gcse Chemistry Exam Questions And Answers
9. Balancing eBooks and Physical Books Gcse Chemistry Exam Questions And Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gcse Chemistry Exam Questions And Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gcse Chemistry Exam Questions And Answers
- Setting Reading Goals Gcse Chemistry Exam Questions And Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gcse Chemistry Exam Questions And Answers
- Fact-Checking eBook Content of Gcse Chemistry Exam Questions And Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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