

1. The theory that explains the behavior of gases at the molecular level is called the kinetic theory. Boyle's law refers to the relationship between the theoretical pressure exerted by an ideal gas.
2. Gases deviate most from ideal gas behavior under conditions of very low temperature and very high pressure. The molecules of an ideal gas display no attraction or repulsion for one another. Under ordinary conditions, an ideal gas consists entirely of single atoms which explain why gases are so easily compressed. Ideal gas particles travel in straight lines until they collide with each other or with the walls of their container. The approximate kinetic energy of an ideal gas is completely proportional. The macroscopic energy of the molecules of an ideal gas is directly proportional to the square temperature of the gas.
3. A gas exerts pressure on the walls of its container because gas molecules collide with the walls of the container. As temperature increases, gas depends on two factors:
 - a. the velocity of the particles that move
 - b. the weight of the particles that move into the surface of the container
4. To measure gas pressure an instrument called a manometer is used.
5. The earth's atmosphere has weight which creates atmospheric pressure.
6. The instrument used to measure atmospheric pressure is the barometer.

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Albert S. Tarendash



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