

100) A student finds the ΔG° of hydrolysis of ATP is -30.5 kJ/mol at pH 7.0. Is this value the same as the ΔG° of hydrolysis of ATP at pH 7.0? Explain.

101) What does the term "coupled reaction" mean?

_____ ΔG° for $\text{H}_2\text{O} \rightarrow \text{H}_2 + \frac{1}{2}\text{O}_2$ is $+237 \text{ kJ/mol}$ _____
 _____ ΔG° for $\text{H}_2 + \frac{1}{2}\text{O}_2 \rightarrow \text{H}_2\text{O}$ is -237 kJ/mol _____

102) What is the standard Gibbs free energy of formation of H_2O at pH 7.0?

_____ ΔG°_f for H_2O is -237 kJ/mol _____

103) Calculate the ΔG° of hydrolysis of ATP at pH 7.0.

$\frac{1.2 \times 10^{-3} \text{ M}}{1.2 \times 10^{-3} \text{ M}} = \frac{1.2 \times 10^{-3} \text{ M}}{1.2 \times 10^{-3} \text{ M}}$ $\Delta G^\circ = -30.5 \text{ kJ/mol}$ $\Delta G = -30.5 \text{ kJ/mol}$
 _____ ΔG° for $\text{H}_2\text{O} \rightarrow \text{H}_2 + \frac{1}{2}\text{O}_2$ is $+237 \text{ kJ/mol}$ _____
 _____ ΔG° for $\text{H}_2 + \frac{1}{2}\text{O}_2 \rightarrow \text{H}_2\text{O}$ is -237 kJ/mol _____

104) What is the ΔG° of hydrolysis of ATP at pH 7.0? Explain.

105) Explain why the ΔG° of hydrolysis of ATP is -30.5 kJ/mol .

_____ The ΔG° of hydrolysis of ATP is -30.5 kJ/mol _____
 _____ The ΔG° of hydrolysis of ATP is -30.5 kJ/mol _____
 _____ The ΔG° of hydrolysis of ATP is -30.5 kJ/mol _____

106) Explain why the ΔG° of hydrolysis of ATP is -30.5 kJ/mol .

_____ The ΔG° of hydrolysis of ATP is -30.5 kJ/mol _____

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