

F321: Atoms, Bonds and Groups

Redox

39 Marks

1. The element strontium forms a nitrate, $\text{Sr}(\text{NO}_3)_2$, which decomposes on heating as shown below.



- (i) Using oxidation numbers, explain why the reaction involves both oxidation and reduction.

.....

.....

.....

.....

[3]

- (ii) A student heats 5.29 g of $\text{Sr}(\text{NO}_3)_2$ and collects the gas at room temperature and pressure, RTP.



Calculate the volume of gas, in dm^3 , obtained by the student at RTP.

Molar mass of $\text{Sr}(\text{NO}_3)_2 = 211.6 \text{ g mol}^{-1}$.

answer = dm^3

[3]

[Total 6 marks]

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L Darling-Hammond



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