

May 2009 TZ1 Question 1

- (b) For part of her extended essay investigation into the efficiency of the process, a student reacted a pure sample of a vegetable oil (where $R = C_{17}H_{33}$) with methanol. The raw data recorded for the reaction is below.

Mass of oil	= 1013.0 g
Mass of methanol	= 200.0 g
Mass of sodium hydroxide	= 3.5 g
Mass of biodiesel produced	= 811.0 g

The relative molecular mass of the oil used by the student is 885.6. Calculate the amount (in moles) of the oil and the methanol used, and hence the amount (in moles) of excess methanol.

[3]

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May 2009 TZ2 Question 1

- (b) A student reacted some salicylic acid with excess ethanoic anhydride. Impure solid aspirin was obtained by filtering the reaction mixture. Pure aspirin was obtained by recrystallization. The following table shows the data recorded by the student.

Mass of salicylic acid used	3.15 ± 0.02 g
Mass of pure aspirin obtained	2.50 ± 0.02 g

- (i) Determine the amount, in mol, of salicylic acid, $C_6H_4(OH)COOH$, used.

[2]

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- (ii) Calculate the theoretical yield, in g, of aspirin, $C_6H_4(OCOCH_3)COOH$.

[2]

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SJ Ball



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web this project involved the simulation of radioactive decay the probability p that a certain atom will decay in a small
amount of time t is $p = \lambda t$ $1 - \lambda t \approx 1 - \lambda t$ $\ln 2 = t_{1/2} \lambda$ where λ

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web the most common types of radioactivity are α decay β decay γ emission positron emission and electron capture nuclear
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