

# C2 - THE PERIODIC TABLE 2.6 The transition metals

## GCSE Chemistry Aiming for Grade 4

Where are the transition metals on the periodic table?  
In the centre block

Give the names of five transition metals

**Gold, Silver, Aluminium, Iron, Copper**

List the physical properties of the transition metals

1. Good conductors of heat
2. Good conductors of electricity
3. High densities
4. High melting points
5. Hard and strong

Why is mercury not typical as a transition metal?

It is a liquid at room temperature  
whereas all the others are solid with  
high melting points.



Assessment  
Grade 4

## GCSE Chemistry Aiming for Grade 6

Compared to Group 1 metals how do the following  
properties compare?

**Density**

Much higher than group 1, none of them will float on  
water like lithium, sodium and potassium.

**Melting point**

Much higher melting points except for mercury

**Boiling point**

Boiling points are extremely high compared to the  
alkali metals

**Hardness**

Alkali metals can be cut with a knife and are soft.  
Transition metals are very hard.

**Strength**

Transition metals are very strong and can be used for  
buildings and machines due to their strength

Give the names of the following transition metal  
compounds.

**Fe<sub>2</sub>O<sub>3</sub>**

**Iron Oxide**

**Cu<sub>2</sub>O**

**Copper Oxide**

**NiCO<sub>3</sub>**

**Nickel Carbonate**

**AgNO<sub>3</sub>**

**Silver Nitrate**

Assessment  
Grade 6

## GCSE Chemistry Aiming for Grade 8

Why is copper used for wires?  
It is a very good conductor -  
other metals. It is also very  
durable into wires.

Steel is made from Iron and  
make girders on bridges?  
It is very strong and hard to  
weight and high forces.

Titanium is used for aircraft  
Titanium has a very high melting  
point ideal for very hot engine parts

State the colours of the compounds  
of transition metals.

Copper sulphate **blue** due to

Nickel Carbonate **pale green**

Chromium oxide **dark green**

Manganese Chloride **pale pink**

Compared with Group 1 elements  
metals have different properties  
Transition metals can form a  
variety of ions. Their reactivity is different.  
eg iron can form Fe<sup>2+</sup> and Fe<sup>3+</sup>  
Different reactivity.

Assessment  
Grade 8

# May 2012 Chemistry C2 Mark Scheme

**R Barnett**



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### **Table of Contents May 2012 Chemistry C2 Mark Scheme**

1. Understanding the eBook May 2012 Chemistry C2 Mark Scheme
  - The Rise of Digital Reading May 2012 Chemistry C2 Mark Scheme
  - Advantages of eBooks Over Traditional Books
2. Identifying May 2012 Chemistry C2 Mark Scheme
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an May 2012 Chemistry C2 Mark Scheme
  - User-Friendly Interface
4. Exploring eBook Recommendations from May 2012 Chemistry C2 Mark Scheme
  - Personalized Recommendations
  - May 2012 Chemistry C2 Mark Scheme User Reviews and Ratings

- May 2012 Chemistry C2 Mark Scheme and Bestseller Lists
- 5. Accessing May 2012 Chemistry C2 Mark Scheme Free and Paid eBooks
  - May 2012 Chemistry C2 Mark Scheme Public Domain eBooks
  - May 2012 Chemistry C2 Mark Scheme eBook Subscription Services
  - May 2012 Chemistry C2 Mark Scheme Budget-Friendly Options
- 6. Navigating May 2012 Chemistry C2 Mark Scheme eBook Formats
  - ePub, PDF, MOBI, and More
  - May 2012 Chemistry C2 Mark Scheme Compatibility with Devices
  - May 2012 Chemistry C2 Mark Scheme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of May 2012 Chemistry C2 Mark Scheme
  - Highlighting and Note-Taking May 2012 Chemistry C2 Mark Scheme
  - Interactive Elements May 2012 Chemistry C2 Mark Scheme
- 8. Staying Engaged with May 2012 Chemistry C2 Mark Scheme
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers May 2012 Chemistry C2 Mark Scheme
- 9. Balancing eBooks and Physical Books May 2012 Chemistry C2 Mark Scheme
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection May 2012 Chemistry C2 Mark Scheme
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine May 2012 Chemistry C2 Mark Scheme
  - Setting Reading Goals May 2012 Chemistry C2 Mark Scheme
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of May 2012 Chemistry C2 Mark Scheme
  - Fact-Checking eBook Content of May 2012 Chemistry C2 Mark Scheme
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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