

## The K<sub>sp</sub> of Lead Iodide, PbI<sub>2</sub>

C12-4-13 & C12-1-01

**Fact:** Many barium compounds are extremely poisonous, yet barium sulfate is routinely swallowed by patients in order to get clearer x-ray pictures of their lower intestines. Why doesn't barium sulfate poison these patients?

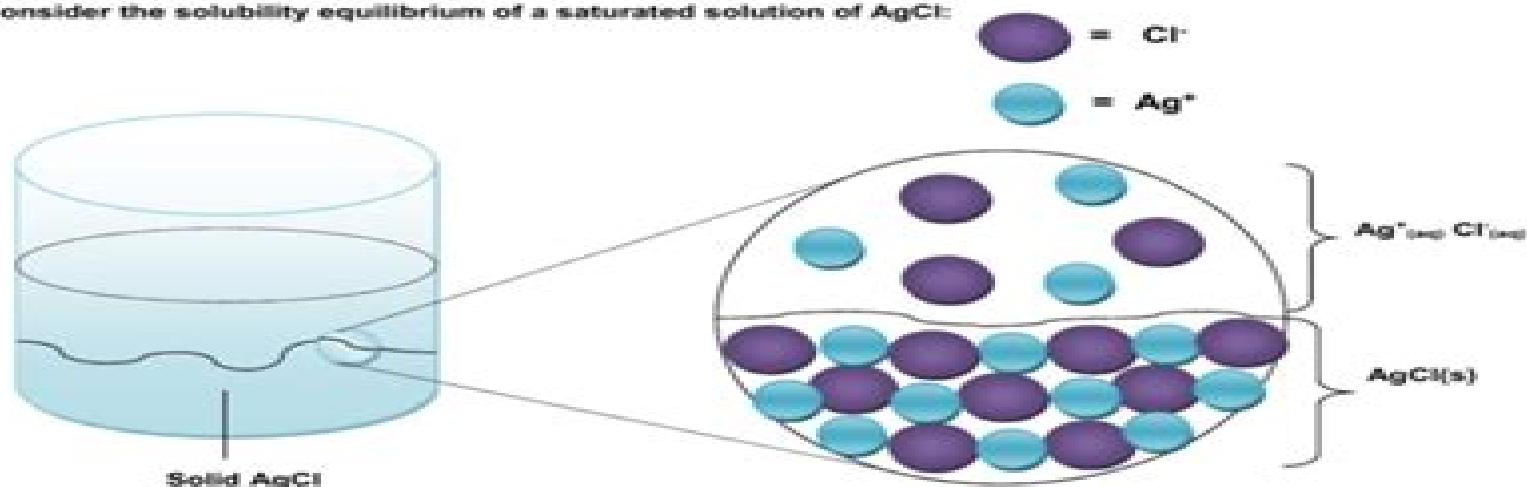
**Fact:** It is possible to saturate a swimming pool with silver bromide and kill off all the nasty little micro-organisms such as algae and bacteria, so that the pool water remains clean and sparkling, yet the water is as pure as drinking water and harmless to people. Since silver ion is a heavy metal ion and hence dangerous to our health in large doses, how is it possible that water saturated with silver bromide doesn't harm us?

The answers to the puzzling questions posed above lies in the low solubility of many salt compounds.

### Introduction:

When ionic solids dissolve, they dissociate into their positive and negative ions that make up the solids. These ions become hydrates and have the same relative proportions when in solution and when solid. The more the solid dissolves, the more the ion concentration increases. This increase in ion concentration allows for the reverse reaction (crystallization) to occur. Eventually the rate of dissolving will equal the rate of crystallization. This state of saturation is called equilibrium. This state can be recognized by a constant color or constant mass.

Consider the solubility equilibrium of a saturated solution of AgCl:



# Ksp Of Lead Iodide Lab Calculations

**R Barnett**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Ksp Of Lead Iodide Lab Calculations:**

**Chemical Principles** Peter Atkins, Loretta Jones, 2007-08 Written for calculus inclusive general chemistry courses Chemical Principles helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications Unlike other texts it begins with a detailed picture of the atom then builds toward chemistry's frontier continually demonstrating how to solve problems think about nature and matter and visualize chemical concepts as working chemists do Flexibility in level is crucial and is largely established through clearly labeling separating in boxes the calculus coverage in the text Instructors have the option of whether to incorporate calculus in the coverage of topics The multimedia integration of Chemical Principles is more deeply established than any other text for this course Through the unique eBook the comprehensive Chemistry Portal Living Graph icons that connect the text to the Web and a complete set of animations students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding *Quantitative Chemical Analysis, Sixth Edition* Daniel C. Harris, 2003 For instructors who wish to focus on practical industrial or research chemistry Includes case studies applications boxes and spreadsheet applications

Chemical Principles in the Laboratory Emil J. Slowinski, 1996 Provides a series of experiments designed to teach students the available experimental methods the proper design of experiments and the interpretation of experimental results

*Experiments in Chemistry* Frank R. Milio, Clyde R. Metz, W. G. Nordulf, 1991-03 *Quantitative Chemical Analysis* Daniel C. Harris, 2010-04-30 QCA is the bestselling textbook of choice for analytical chemistry It offers a modern portrait of the techniques of chemical analysis backed by a wealth of real world applications This edition features new coverage of spectroscopy and statistics new pedagogy and enhanced lecturer support **Exploring Chemical Analysis** Daniel C. Harris, 2005 Exploring Chemical Analysis teaches students how to understand analytical results and how to use quantitative manipulations preparing them for the problems they will encounter **Chemical Principles in Calculations of Ionic**

**Equilibria** Emil J. Margolis, 1966 **CHEMISTRY EXPERIMENTS** James Signorelli, 2014-09-19 Gifted and talented students and any student interested in pursuing a science major in college needs a rigorous program to prepare them while they are still in high school This book utilizes a format where the application of several disciplines science math and language arts principles are mandated Each lab concludes with either an essay or a detailed analysis of what happened and why it happened This format is based on the expectations of joining a university program or becoming an industrial science professional The ideal student lab report would be written in a lab research notebook and then the essay or final analysis is done on a word processor to allow for repeat editing and corrections The research notebook has all graph pages a title section and a place for the students and their assistants to sign and witness that exercise The basic mechanics of the lab report title purpose procedure diagrams data table math and calculations observations and graphs are handwritten into the book The conclusion is done on a word processor MS Word which allows the instructor to guide the student in writing and

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**Principles of Modern Chemistry** David W. Oxtoby,H. P. Gillis,Norman H. Nachtrieb,2002 This successful textbook continues to address students in honours and high mainstream general chemistry courses The chapters in the fourth edition are organised within larger units to emphasise the conceptual structure of chemistry     Biology ,1993     *(Free Sample)*  
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**Chemistry & Chemical Reactivity** John C. Kotz,Paul Treichel,Patrick A. Harman,2003 Provides a broad overview of the principles of chemistry the reactivity of chemical elements and their compounds and the applications of chemistry Conveys a sense of chemistry as a field that not only has a lively history but also one that is currently dynamic with important new developments on the horizon     **Modern Experiments for Introductory College Chemistry** Howard A. Neidig,William

Franklin Kieffer,1967     **Russian Journal of Physical Chemistry** ,1984     **Fundamental Aspects of Crystallization and Precipitation Processes** Daina M. Briedis,1987     **New Trends in Chemistry Teaching** Edward Cartmell,1967

AIChE Symposium Series American Institute of Chemical Engineers,1986     Lead Iodide X-ray and Gamma-ray Spectrometers for Room and High Temperature Operation ,1997 In this study we report on the results of the investigation of lead iodide material properties The effectiveness of zone refining purification methods on the material purity is determined by ICP MS and ICP OES and correlated to the electrical and physical material properties We show that this zone refining

method is very efficient in removing impurities from lead iodide and we also determine the segregation coefficient for some of these impurities Triple axis x ray diffraction TAD analysis has been used to determine the crystalline perfection of the lead iodide after applying various cutting etching and fabrication methods The soft lead iodide crystal was found to be damaged when cleaved by a razor blade but by using a diamond wheel saw followed by etching the crystallinity of the material was improved as observed by TAD Low temperature photoluminescence also indicates an improvement in the material properties of the purified lead iodide Electrical properties of lead iodide such as carrier mobility were calculated based on carrier phonon scattering The results for the electrical properties were in good agreement with the experimental data

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