In a chemical reaction <u>atoms</u> combine in different ways to form new <u>products</u>.
 Whatever the change, you cannot <u>lose</u> any atoms. This is known as the law of <u>conservation</u> of

reaction.

- mass. It means that if you could weigh the <u>reactants</u> at the start, and then compare this with the <u>mass</u> of all the <u>products</u> formed at the end, they would be <u>equal</u>.

  3a. When a metal is heated in air it <u>gains</u> mass in forming the metal oxide.

  9Hwf4ans
- 3b. When a carbonate dissolves in acid, the mixture <u>loses</u> mass.

  3c. Both these situations involve the 'adding' or losing' of a gas, which is rather difficult to weigh!

  4. From the 'law' in 2. you can make <u>predictions</u> about the quantities of chemicals involved in a
- 4a. For example, 24g of magnesium burns in air and combines with 16g of oxygen to form 40g of magnesium <u>oxide</u>, even if you can't weigh the oxygen!
- 4b. This means 36g of magnesium will combine with 24g of oxygen from the air to form 🕰 of magnesium oxide.
- 4c. If g of magnesium is burned it combines with 4g of oxygen to form 10g of magnesium oxide.

  4d. Burning 12g of magnesium will combine with g of oxygen to make g of magnesium oxide
- 4e. burning 40g of magnesium will combine with 32g of oxygen to make 80g of magnesium oxide.

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