Draw the following substances. Then, identify <u>the strongest</u> intermolecular force present in pure samples of the following substances:

SO ₂	H ₂ O	CH ₂ Cl ₂
**	TO TO	· · · · · · · · · · · · · · · · · · ·
dipole-dipole forces	H H	H CF dipole-dipole forces
sco	PCI ₃	SO_3
Ö::C::S	:cí:gi;ci:	

Identify the strongest intermolecular force operating in the condensed phases of the following substances. Fully explain how you determined this.

London dispersion forces

a. Cl ₂ London dispersion forces	b. CO Dipole-dipole forces
The CI-CI bond is nonpolar so the molecule is nonpolar. Non polar molecules have only London dispersion forces operating in the substance.	The C-O bond is polar so the molecule is polar. Polar molecules have dipole-dipole forces. They also have London dispersion forces, but dipole-dipole forces are stronger.
e. SO ₂ Dipole-dipole forces	d. CH ₂ Cl ₂ Dipole-dipole forces
SO ₂ is a bent, polar molecule. The strongest intermolecular force in a polar molecule is the dipole-dipole force	The strongest intermolecular force in a polar molecule that cannot form hydrogen bonds is the dipole-dipole force
e. HF Hydrogen bonding forces	g. CH ₃ -O-CH ₃ Dipole-dipole forces
Molecules that have hydrogen attached to an O, N, or F can form hydrogen bonds. These are the strongest of the intermolecular forces.	The hydrogen atoms are not bonded to the oxygen, so this molecule cannot form hydrogen bonds. It is polar, so it will have dipole-dipole forces.

dipole-dipole forces

dipole-dipole forces

<u>Intermolecular Forces And Organics Worksheet Answer</u> <u>Key</u>

Clifford Lane

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