CHAPTER TEST

Class 11 - Chemistry

1. (a) Propene

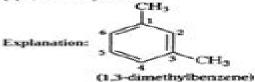
Explanation: Elimination reaction takes place, a molecule of HBr gets eliminated to form propene.

2.

(c) HI > HBr > HCI

Explanation: HI > HBr > HCI

3. (a) 1,3-dimethy/fbenzene



4.

(d) Geometric isomers

Explanation: Isomers which have the same structural formulae but differ in the relative spatial arrangements of atoms or groups around the double bond are called geometric isomers. Thus geometrical isomers are a type of stereoisomers.

(d) HgSO₄

Explanation: Alkynes are not readily hydrated due to their low reactivity towards electrophilic addition reactions. Further dilute H₂SO₄ has no effect on carbon carbon triple bond. In the presence of HgSO4 as catalyst, hydration occurs readily.

6. (a) A mixture of low boiling hydrocarbons.

Explanation: Pyrolysis is the decomposition of a compound by the application of heat. It involves the breaking of C-C and C-H bonds resulting in the formation of lower boiling alkanes.

7. (a) 3

Explanation: The position isomers of C5H₈ are n-pentyne, 3-methyl butyne, pent-2-yne.

(a) C_nH_{2n}

Explanation: Alkenes are unsaturated hydrocarbons with at least on double bond between the carbon atoms. Their general formula is C_nH_{2n} .

9.

(b) Sodium acetate

Explanation:

This is an example of Kolbe's electrolysis method. The reaction is:

$$2CH_3COONa + 2H_2O \xrightarrow{\text{obst-dipost}} CH_3$$
. $CH_3 + 2NaOH + H_2 + 2CO_2$

The step-wise redox reactions occuring in the electrolytic cell are depicted as under

$$\begin{array}{c} \text{Electrolysis} \\ 2\text{CH}_2\text{COO} \, \text{Nu} + 2\text{H}_2\text{O} & \text{CH}_3 - \text{CH}_3 + 2\text{CO}_2 + \text{H}_2 + 2\text{NuOH} \\ \text{At Anodec (Oxidation)} \\ 2\text{CH}_2 - \frac{1}{6} - 0^- - 2e^- - 2\text{CH}_3 & \frac{1}{6} - 0 & -2\text{CH}_3 + 2\text{CO}_2 \\ 2\text{CH}_3 - \text{CH}_3 - \text{CH}_3 & \frac{1}{6} - 0 & \frac{1}{6} - \frac{1}{6} & \frac{1}{6} - \frac{1}{6} & \frac{1}{6}$$

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