Question 33 (4 marks)

A student adds 1.17 g of Al(OH)₃(s) to 0.500 L of 0.100 mol L⁻¹ HCl(aq).

Calculate the pH of the resulting solution. Assume that the volume of the resulting solution is 0.500 L.

1) Write a balanced chemical equation

A1 (0H)3 (5) + 3 HC (109) - A1 C13 (49 1+ 3 H2 OC)

(2) Find no. of mole > of QH and H+ $n(AI(OH)_3) = \frac{m}{mm} = \frac{1}{24.98 + 3(16 + 1.008)} = 0.01499.$

 $n(0H^{-}) = 3 \times n(AI(0H)_{3}) = 0.04499...mp1$

 $N(H(I) = C \times V = 0.5 \times 0.1 = 0.05 \text{ mol}$

 $n(H^{+}) = n(HCI) = 0.05 mel$

3) Determine which reagent is in excess and now much

D. 05 - 0.04499 .. = 0.00500 ... MD)

1 at 1 the [H+] and thus find pH.

PH = -109 10 CH+7 = 2.000 (3 dp)

Please turn over

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Table of Contents Hsc Chemistry 6 User Guide

- 1. Understanding the eBook Hsc Chemistry 6 User Guide
 - The Rise of Digital Reading Hsc Chemistry 6 User Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Hsc Chemistry 6 User Guide
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc Chemistry 6 User Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Hsc Chemistry 6 User Guide
 - Personalized Recommendations
 - Hsc Chemistry 6 User Guide User Reviews and Ratings
 - Hsc Chemistry 6 User Guide and Bestseller Lists
- 5. Accessing Hsc Chemistry 6 User Guide Free and Paid eBooks
 - Hsc Chemistry 6 User Guide Public Domain eBooks
 - Hsc Chemistry 6 User Guide eBook Subscription Services
 - Hsc Chemistry 6 User Guide Budget-Friendly Options

- 6. Navigating Hsc Chemistry 6 User Guide eBook Formats
 - o ePub, PDF, MOBI, and More
 - Hsc Chemistry 6 User Guide Compatibility with Devices
 - Hsc Chemistry 6 User Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Hsc Chemistry 6 User Guide
 - Highlighting and Note-Taking Hsc Chemistry 6 User Guide
 - Interactive Elements Hsc Chemistry 6 User Guide
- 8. Staying Engaged with Hsc Chemistry 6 User Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc Chemistry 6 User Guide
- 9. Balancing eBooks and Physical Books Hsc Chemistry 6 User Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc Chemistry 6 User Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Chemistry 6 User Guide
 - Setting Reading Goals Hsc Chemistry 6 User Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Chemistry 6 User Guide
 - Fact-Checking eBook Content of Hsc Chemistry 6 User Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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