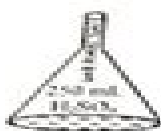


# অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার

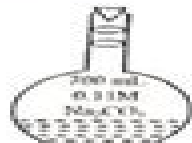


পট্টাবলী কপূরা, এখানে মোট পট্টাবলী, আরও কয়েকটি নির্দেশাবলী কয়েকটি পট্টাবলী এবং মনে রাখতে হবে যে পট্টাবলী কপূর উত্তর অধ্যায়ভিত্তিক মেসেজ রয়েছে। এগুলোর অনুশীলন করলে তুমি এ অধ্যায়টি থেকে সেরাভাবে বুঝতে পারবে।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০/১ এর নং ১/

- মোলের প্রকণ কী?
- "কেলাস পানিভূত লবণের দ্রাব্যতা, 'আপনাত্মা বৃষ্টির সাথে সাথে এক পর্যায়ে বৃষ্টি না পেয়ে হ্রাস পায়'— ব্যাখ্যা করো।
- A-পাত্রে বিন্যাসন এলিটের ভর কত নির্ণয় করো।
- দ্রব দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা পরিষ্কৃত হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

১০০০ mL দ্রবণে কোন দ্রবের 1 mol দ্রবীভূত থাকলে উক্ত দ্রবকে মোলার দ্রবণ বলে।

কেলাস পানিভূত লবণের দ্রাব্যতা আপনাত্মা বৃষ্টির সাথে সাথে এক পর্যায়ে বৃষ্টি না পেয়ে হ্রাস পায়। এর কারণ, আপনাত্মা বৃষ্টির সাথে সাথে দ্রাব্যতা বৃষ্টি পেতে থাকলেও একটি নির্দিষ্ট আপনাত্মা বৃষ্টির পানি দ্রব লবণ নিষ্কাশিত হয়। তখন দ্রাব্যতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর কথা বলা যায়।  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর ক্ষেত্রে পানিতে দ্রবীভূত করার সময় বিয়োজন প্রক্রিয়াটি আপনাত্মা বৃষ্টি 32.4°C আপনাত্মার উপরে  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  যখন  $\text{Na}_2\text{SO}_4$  এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি আপোৎপাদী হয়। এ কারণে 32.4°C আপনাত্মার উপরে আপনাত্মা বৃষ্টিজনিত কারণে দ্রাব্যতা হ্রাস ঘটে।

২

আমরা জানি,  
 $w = SMV$   
 $= (0.5 \times 98 \times 0.25) \text{ g}$   
 $= 12.25 \text{ g}$

অতএব, দ্রবীভূত  $\text{H}_2\text{SO}_4$  এর ভর  
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রে অবস্থিত,  
 $\text{H}_2\text{SO}_4$  এর আণবিক ভর,  
 $M = 1 \times 2 + 32 + 16 \times 4$   
 $= 2 + 32 + 64$   
 $= 98 \text{ g/mol}$

দ্রবণের ঘনমাত্রা,  
 $S = 0.5 \text{ M}$   
 দ্রবণের আয়তন,  
 $V = 250 \text{ mL}$   
 $= \frac{250}{1000} = 0.25 \text{ L}$

দ্রবীভূত  $\text{H}_2\text{SO}_4$  এর ভর,  $w = ?$

১০ পাত্র-A:  $\text{H}_2\text{SO}_4$  এর মোল সংখ্যা  $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$   
 $= 0.125 \text{ mol}$

পাত্র-B:  $\text{Na}_2\text{CO}_3$  এর মোল সংখ্যা  $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$   
 $= 0.077 \text{ mol}$

সম্প্রতিক বিক্রিয়া:  $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$

সমীকরণ হতে পাই, 1 mol  $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$   
 1 mol  $\text{Na}_2\text{CO}_3$  বিক্রিয়া করে  $\text{H}_2\text{SO}_4$  এর 1 mol এর সাথে  
 $0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট  $\text{H}_2\text{SO}_4$  এর পরিমাণ  $= (0.125 - 0.077) \text{ mol}$   
 $= 0.048 \text{ mol}$

পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে 0.048 mol  $\text{H}_2\text{SO}_4$  অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লীয়।

প্রশ্ন ২ একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর তরুণ পরিমাণ এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2%  $\frac{W}{V}$  HCl এর নমুনা কাল করে।

সি. কে. ১০/১ এর নং ১/

- নোড কাকে বলে?
- $\text{Cl}_2$  অপোলার— ব্যাখ্যা করো।
- উদীপকের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

১. নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

২. সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা এক হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হবে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু Cl এর তড়িৎ ঋণাত্মকতা 3.0 ও C অপুতে তড়িৎ ঋণাত্মকতার পার্থক্য  $(3 - 3) = 0$ , তাই  $\text{Cl}_2$  বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এহুণ ঘ থাকে।

৩. উদীপকে বর্ণিত এলিটটি হলো 2%  $\left(\frac{W}{V}\right) \text{ HCl}$

2%  $\left(\frac{W}{V}\right) \text{ HCl}$  এর অর্থ হলো—

100 mL দ্রবণে HCl এর ভর 2g

∴ 1000 mL " " " "  $\left(\frac{2 \times 1000}{100}\right) \text{ g}$   
 $= 20 \text{ g}$

# Hsc Chemistry 1st Paper Question 2solution

**Will Marchment**



## **Hsc Chemistry 1st Paper Question 2solution:**

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whether they re mating for life or just for the breeding season each species has evolved its own recipe for sexual success

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