পরীকারী কপুরা, এখানে বোর্র পরীকা, ক্যাডোট কলেলগর শীর্ষান্ধানীর কলেলগমূহের পরীক্ষা এবং মহেল টেন্টের ব্যাত্তর পূর্ণাল উত্তর অধ্যায়ভিত্তিক দেবছা ময়েছে। এগুলো অনুশীলন করনে কৃষি এ অধ্যায়টি লেকে সেকোনো স্থানশীল জনামূদ্

## 





0.000

100 100 30 B 400 40 31

- ক, মোলার দ্রবণ কীয়
- "কেলাস পানিত্র লকণের দ্রাকারা, তাপমারা বৃশ্বির সাথে সাথে এক পরীয়ে বৃশ্বি না পেয়ে য়াস পায়"— ব্যাব্যা করে।
- চবশ দৃটি মিশ্রিত করলে, থিপ্রদটি কী পরিমাণ অগ্নীয় বা ফারীয় মবে— বিশ্লেষণ করো।

#### ১ নং প্রয়ের উত্তর

1000 mL চৰণে কোন দ্ৰবের 1 mol দ্রবীভূত থাকলে উপ্ত দ্রবণকে
মোলার দ্রবন বলে।

ত্রলাস পানিত্র গবণের প্রাব্যান্তা তাপমাত্রা বৃশ্বির সাথে সাথে এক পর্যান্ত বৃশ্বির না পেরা ছাস পার। এর কারণ, তাপমাত্রা বৃশ্বির সাথে সাবে প্রব্যান্ত বৃশ্বির বিধান পার। এর কারণ, তাপমাত্রা বৃশ্বির সাথে সাবে প্রব্যান বৃশ্বির পোনে থাকার একটি নির্নিন্দ তাপমাত্রার কেলাস পানি কুর গরুণ নির্নান্ত হয়। তথান প্রাব্যান্ত হাস পোনে থাকে। উদাহরণমূল্প Na<sub>2</sub>SO<sub>4</sub>.10H<sub>2</sub>O এর কথা করা যায়। Na<sub>2</sub>SO<sub>4</sub>.10H<sub>2</sub>O এর পেরে পানিতে প্রবীত্ত হওয়ার সময় বিয়োজন প্রক্রিয়াটি ভাপহারী বিক্রিয়া কিন্তু 32.4°C তাপমাত্রার উপরে Na<sub>2</sub>SO<sub>4</sub>.10H<sub>2</sub>O যখন Na<sub>2</sub>SO<sub>4</sub> এ পরিণত হয় তথান বিয়োজন প্রক্রিয়াটি ভাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা কৃশ্বিজনিত কারণে প্রাব্যান্ত হস হটে।

#### 191

भ = SMV

-(0.5 × 98 × 0.25) g

= 12.25 g

অতএব, দ্বীভূত H₂SO₄ এর ভর ≒ 12.25 g

উদ্দীপকের A-পাত্রে অবস্থিত,

H<sub>2</sub>SO<sub>4</sub> এর আণবিক ভর,

M = 1 × 2 + 32 + 16 × 4

= 2 + 32 + 64

= 98 g/mol

রবণের ঘনমারা,

S = 0.5 M

রবণের আরতন,

V = 250 mL

= 250

1000 = 0.25L

ফরীভূত H<sub>2</sub>SO<sub>4</sub> এর ভর, w = 1

ভারত ন: H<sub>2</sub>SO<sub>4</sub> এর মোল সংখ্যা —  $\left(0.5 \times \frac{250}{1000}\right)$  med — 0.125 mod — 0.125 mod — 0.125 mod — 0.077 mod — 0.077 mod — 0.077 mod সংঘটিত বিভিন্ন: H<sub>2</sub>SO<sub>4</sub> + Na<sub>2</sub>CO<sub>3</sub> — 0.077 mod — 0.077

পাত্র-∧ এর প্রবংগর সাথে পাত্র-৪ এর প্রবংগর বিভিন্নার ফলে ১১৯ moi H₂SO₄ অবশিন্ট থাকায় মিশ্রিত প্রবংগর প্রকৃতি মধে অঞ্জীত

19.05

ett.

ভটা ১২ একজন ছাত্র লাবরেটরিতে X-পশ্বতিতে 10 mg বর জেছ হ পরিমাণ এবং Y-পশ্বতিতে 0.5 g পরিমাণ 2% ₩ HCl এর নমূর হ কাজ করে।

- ক, নোড কাকে বলে?
- খ. Cl<sub>2</sub> অপোলার- ব্যাখ্যা করে।।
- উদ্দীপকের এসিভটির মোলার ঘনমারা নির্ণয় করে।
- য়, উদ্দীপকের কোন পশ্বতি পরিবেশবান্ধর? বিশ্লেষণ করে। ২ নং প্রয়োর উত্তর

নিউরিয়াসের চারনিকে যে নির্দিউ ত্রিমাত্রিক স্থানে কোন ইত্রু পাওয়ার সম্ভাব্যতা শূন্য তাকে নোভ বলে।

সমযোজী বন্ধনে আবন্ধ উভয় পরমাণুর তড়িং কণাছকতা ক্র হলে অর্থাৎ তড়িং কণাব্যকতার পার্থকা শূন্য হলে সমযোজী ক বিশুন্ধ ও অপোলার হয়। আর তড়িং ঝণাব্যকতার পার্থকা 0.5 এর হ মর্লে অপোলার হয়। যেহেত CI এর তড়িং ঝণাব্যকতা 3.0 ৫ ট অপুতে তড়িং ঝণাব্যকতার পার্থকা (3 – 3)= 0, তাই CI<sub>2</sub> ক্রিন্ সমযোজী ও অপোলার যৌগ। উরোধ্য মৌলিক অপুর ক্রেত্রে এর্ণ হা থাকে।

ত্রী উদ্দীপকে বর্ণিত এগিচটি হলো 2%  $\left(\frac{W}{V}\right)$  HCI ।

2%  $\left(\frac{W}{V}\right)$  HCI এর অর্থ ফলো-

.. 1000 mL ... ... ...  $\frac{2 \times 1000}{100}$  g

# **Hsc Board Chemistry Answer Paper**

Ying Liu

#### **Hsc Board Chemistry Answer Paper:**

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## **Table of Contents Hsc Board Chemistry Answer Paper**

- 1. Understanding the eBook Hsc Board Chemistry Answer Paper
  - The Rise of Digital Reading Hsc Board Chemistry Answer Paper
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Hsc Board Chemistry Answer Paper
  - Exploring Different Genres
  - o Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Hsc Board Chemistry Answer Paper
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Hsc Board Chemistry Answer Paper
  - Personalized Recommendations
  - Hsc Board Chemistry Answer Paper User Reviews and Ratings

- Hsc Board Chemistry Answer Paper and Bestseller Lists
- 5. Accessing Hsc Board Chemistry Answer Paper Free and Paid eBooks
  - Hsc Board Chemistry Answer Paper Public Domain eBooks
  - Hsc Board Chemistry Answer Paper eBook Subscription Services
  - Hsc Board Chemistry Answer Paper Budget-Friendly Options
- 6. Navigating Hsc Board Chemistry Answer Paper eBook Formats
  - o ePub, PDF, MOBI, and More
  - Hsc Board Chemistry Answer Paper Compatibility with Devices
  - Hsc Board Chemistry Answer Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Hsc Board Chemistry Answer Paper
  - Highlighting and Note-Taking Hsc Board Chemistry Answer Paper
  - Interactive Elements Hsc Board Chemistry Answer Paper
- 8. Staying Engaged with Hsc Board Chemistry Answer Paper
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Hsc Board Chemistry Answer Paper
- 9. Balancing eBooks and Physical Books Hsc Board Chemistry Answer Paper
  - $\circ$  Benefits of a Digital Library
  - Creating a Diverse Reading Collection Hsc Board Chemistry Answer Paper
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc Board Chemistry Answer Paper
  - Setting Reading Goals Hsc Board Chemistry Answer Paper
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Hsc Board Chemistry Answer Paper
  - Fact-Checking eBook Content of Hsc Board Chemistry Answer Paper
  - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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