



# EMC® Unisphere for VMAX

VERSION 1.5

## Unisphere for VMAX Product Guide

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REV 01

# Emc Unisphere Lab Guide

**L. K. Maheshwari, M. M. S. Anand**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the authors' names.

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is  $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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1  $\text{K}_3\text{PO}_4$  aq + 1  $\text{Na}_3\text{PO}_4$  s  $\rightarrow$  3  $\text{KNO}_3$  aq + 3  $\text{NaNO}_3$  aq net ionic  $\text{PO}_4^{3-}$  aq +  $\text{PO}_4^{3-}$  s  $\rightarrow$  3  $\text{K}^+$  aq + 3  $\text{Na}^+$  aq

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13  $2\text{ag aq} + 2\text{f aq} + 2\text{nh}_4\text{ aq} + \text{cr}_2\text{o}_7^{2-}\text{ aq} \rightarrow \text{ag}_2\text{cr}_2\text{o}_7\text{ s} + 2\text{nh}_4\text{ aq} + 2\text{f aq}$  because two nh 4 aq

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