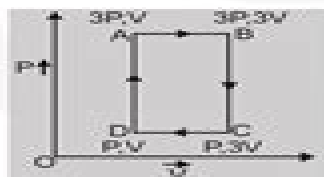


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- Pressure-temperature relationship for an ideal gas undergoing adiabatic change is ( $\gamma = C_p/C_v$ )  
 (1)  $pT^\gamma = \text{constant}$  [AIPMT 92, HARYANA PMT 01, RPET 2001]  
 (2)  $pT^{1-\gamma} = \text{constant}$   
 (3)  $pT^{1+\gamma} = \text{constant}$   
 (4)  $p^{1-\gamma}T^\gamma = \text{constant}$
- An inventor claims that when temperatures of source and sink are  $127^\circ\text{C}$  and  $27^\circ\text{C}$  respectively then efficiency of engine is 26%, then :- [AIPMT 2001]  
 (1) It is impossible (2) It is possible but possibility is less  
 (3) It's possibility is high (4) Data is insufficient
- If the system takes 100 cal. heat, and releases 80 cal to sink, if source temperature is  $127^\circ\text{C}$  find the sink temp. :- [AIIMS 2001]  
 (1)  $47^\circ\text{C}$  (2)  $127^\circ\text{C}$  (3)  $67^\circ\text{C}$  (4)  $107^\circ\text{C}$
- The efficiency of carnot engine is 50% and temperature of sink is 500K. If temperature of source is kept constant and its efficiency raised to 60%, then the required temperature of the sink will be:- [MANIPAL 2001, AIPMT 2002]  
 (1) 100 K (2) 600 K (3) 400 K (4) 500 K
- An ideal gas heat engine operates in carnot cycle between  $227^\circ\text{C}$  and  $127^\circ\text{C}$ . It absorbs  $6 \times 10^4$  cal of heat at higher temperature. Amount of heat converted to work is. [AIPMT 2005]  
 (1)  $2.4 \times 10^4$  cal (2)  $6 \times 10^4$  cal (3)  $1.2 \times 10^4$  cal (4)  $4.8 \times 10^4$  cal
- An ideal monoatomic gas is taken round the cycle ABCDA as shown in following P - V diagram. The work done during the cycle is : [RPET 92, 2000, IIT 83, AIPMT 90, 92, 2007]



- (1) PV (2) 2 PV (3) 4 PV (4) Zero
- The  $\left(\frac{W}{Q}\right)$  of a carnot-engine is  $\frac{1}{6}$ , now the temp. of sink is reduced by  $62^\circ\text{C}$ , then this ratio becomes twice, therefore the initial temp. of the sink and source are respectively :- [AIPMT 2007]  
 (1)  $33^\circ\text{C}$ ,  $67^\circ\text{C}$  (2)  $37^\circ\text{C}$ ,  $99^\circ\text{C}$   
 (3)  $67^\circ\text{C}$ ,  $33^\circ\text{C}$  (4) 97 K, 37 K

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**S Marginson**



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