



Tech. Degree I & II Semester (Combined) Examination, June 2008

IT/CS/EC/CE/ME/SE/EB/EI/EE/FT 103 ENGINEERING CHEMISTRY
(2006 Scheme)

3 Hours

Maximum Marks: 100

PART A

(Answer all questions)

(8 x 5 = 40)

- | | | |
|---|--|--|
| I | ☒ a) Explain Born-Haber cycle with example.
☐ b) Explain in brief what is BET isotherm.
☐ c) What is photovoltaic effect? Give details of solar cells.
☐ d) Explain the significance of Arrhenius equation.
☐ e) What is Trouton's rule? Explain.
☒ f) What is meant by chemical potential? Give its importance in chemical reactions.
☒ g) Write notes on optical fibres.
☐ h) Write notes on nanoceramics. | |
|---|--|--|

PART B

(15 x 4 = 60)

- | | | |
|--------|--|-------------------|
| II ✓ | ☐ a) Explain in details Point Defects.
☐ b) Explain Band theory of solids. | (10)
(5) |
| OR | | |
| III ✓ | ☐ a) Explain the different types of liquid crystals.
☐ b) Write notes on Super Conductors. | (10)
(5) |
| IV | ☐ a) Explain the Poggendorff compensation method for determination of emf.
☐ b) What is Buffer solution? Explain its action with examples.
☐ c) Explain Cathodic protection method in metallic corrosion. | (6)
(4)
(5) |
| OR | | |
| V ✓ | ☒ a) Give the important factors controlling corrosion.
☐ b) What is a sec-cell? Explain the working of lead-acid storage cell. | (8)
(7) |
| VI ✓ | ☒ a) Derive the Kirchhoff's equation.
☐ b) Explain Nernst heat theorem.
☐ c) Thermodynamics of Biochemical reactions. | (6)
(4)
(5) |
| OR | | |
| VII ✓ | ☐ a) Derive the Gibbs-Helmholtz equation.
☐ b) State the three different rules of thermodynamics.
☐ c) Show and explain the relation between temperature, pressure and free energy. | (5)
(6)
(4) |
| VIII ✓ | ☐ a) What is polymer processing? Explain in details the different methods.
☐ b) Explain the mechanism of lubrication. | (10)
(5) |
| OR | | |
| IX | ☐ a) Write down in detail the preparation, properties and application of any three industrial polymers.
☒ b) What are the different types of lubricants? Give the details. | (10)
(5) |

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