

Mark Scheme

Q1.

| Question Number | Acceptable Answer | Additional Guidance | Mark |
|-----------------|--|---|------|
| (a) (i) | $\text{CaCO}_3(\text{s}) + 2\text{HCl}(\text{aq}) \rightarrow \text{CaCl}_2(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$ <p>Balanced equation (1)</p> <p>State symbols (1)</p> | <p>Accept $\text{CaCO}_3(\text{s}) + 2\text{H}^+(\text{aq}) \rightarrow \text{Ca}^{2+}(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$</p> <p>2nd mark dependent on first or near miss.</p> <p>Reject $\text{H}_2\text{CO}_3(\text{aq})$ in equation, but allow state symbol mark if otherwise correct.</p> | (2) |
| Question Number | Acceptable Answer | Additional Guidance | Mark |
| (a) (ii) | Finds molar mass of calcium carbonate | <p>Example of calculation $\text{Mr of calcium carbonate} = 40.1 + 12 + (16 \times 3) = 100.1 \text{ (g mol}^{-1}\text{)}$ Allow $= 40 + 12 + (16 \times 3) = 100 \text{ (g mol}^{-1}\text{)}$ Accept answer with no working</p> | (1) |

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Shasha Hu



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