Section A.

Answer all questions in the spaces provided

Anhydrous magnesium chloride, MgCl₂, can absorb water to form the hydrated salt 0 1 MarCla,4HaO

Suggest one reason why the enthalpy change for this reaction cannot be determined directly by calorimetry.

[1 mark]

Virtually impossible to prevent dissolving.

Some enthalpies of solution are shown in Table 1. 0 1 -

Table 1

making the salts

Salt	Enthalpy of solution / kJ mol ⁻¹
MgCl ₂ (s)	-155
MgCl ₂ .4H ₂ O(s)	-39

Calculate the enthalpy change for the absorption of water by MgCl₂(s) to form $MgCl_2.4H_2O(s).$

[2 marks]

kJ mol⁻¹ Enthalpy change



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D Kirk

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