Answer Key.

Calculating Specific Heat Worksheet

Name:		Det.
		100

 $Q = mc\Delta T$, where Q = heat energy, m = mass, and $\Delta T = change in temp. Remember, <math>\Delta T = (T_{treat} - T_{initial})$. Show all work and proper units.



 A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

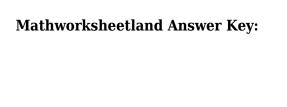
 How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.

 100.0 g of 4.0°C water is heated until its temperature is 57°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.

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Victor M. Corman



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