



CHEMISTRY SSC-II

2

Time allowed: 2:40 Hours

Total Marks Sections B and C: 53

NOTE:- Answer any eleven parts from Section 'B' and any two questions from Section 'C' on the separately provided answer book. Use supplementary answer sheet i.e. Sheet-B if required. Write your answers neatly and legibly.

SECTION – B (Marks 33)

Q. 2 Attempt any ELEVEN parts. The answer to each part should not exceed 3 to 4 lines. (11 x 3 = 33)

- (i) a. Why is the boiling point of Water higher than that of H_2S and H_2Se ?
b. Write down the uses of Heavy water.
- (ii) Complete the following chemical equations and state conditions:
a. $Na + H_2O \rightarrow$ b. $Zn + H_2O_{(g)} \rightarrow$ c. $C + H_2O_{(g)} \rightarrow$
- (iii) Write the chemical reaction of Carbon dioxide with Water and Carbon.
- (iv) a. Name the crystalline forms of Silicon dioxide?
b. Draw the structure of Silica.
- (v) In the laboratory oxygen is prepared by heating a mixture of Potassium Chlorate and Manganese dioxide.
a. Write down the balanced chemical equation of the above reaction.
b. What is the role of Manganese dioxide in this reaction?
c. Write one difference between O_2 and O_3 .
- (vi) a. Write the reaction of Ammonia with Water.
b. Write any four uses of Ammonia.
- (vii) Compare Rhombic Sulphur with Monoclinic Sulphur.
- (viii) a. Write two chemical reactions proving that H_2SO_4 is a strong oxidizing agent.
b. What is Vulcanization?
- (ix) Write the chemical reactions of Chlorine with the following:
a. Sb b. NH_3 c. CO
- (x) Write down the Laboratory and Industrial preparation of Bleaching powder.
- (xi) Write down the properties of Alloys.
- (xii) Write Goldschmidt's Thermite process.
- (xiii) a. Define Fractional distillation. Name the hydrocarbons obtained as a result of fractional distillation of petroleum.
b. Write about the refining of petroleum in Pakistan.
- (xiv) a. Define Isomerism.
b. Write two structural formulae of isomers of Butane. How do these isomers differ from each other?
- (xv) a. What is Soap? Also write the raw material for the manufacture of Soap.
b. What is mixed with kitchen soap?

SECTION – C (Marks 20)

Note: Attempt any TWO questions. All questions carry equal marks. (2 x 10 = 20)

- Q 3.** a. Define Allotropy. Explain the amorphous forms of Carbon. (07)
b. Write any three differences between Diamond and Graphite. (03)
- Q 4.** a. Define Oxides. Explain the types of oxides on the basis of valence number. (07)
b. Explain Chlorine as a Bleaching agent and Highly reactive agent. (03)
- Q 5.** a. Write the laboratory preparation and uses of Methane. (2+2)
b. Write the chemical reactions of Methane with the following. (04)
 (i) Steam (ii) Concentrated Nitric Acid
 (iii) Halogens (iv) Oxygen
 c. Write down the advantages of Open Hearth process in steel manufacturing. (02)

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