

সিলেট বোর্ড

বহুনির্বাচনি অধীক্ষা

তদারকন

প্রথম পত্র

[২০২৩ সালের সিলেটবাস অধ্যায়ী]

সময়—২৫ মিনিট

পূর্ণমান—২৫

বিষয় কোড :

1 7 6

সেট :

খ

[বিশেষ দ্রষ্টব্য : সরবরাহকৃত বহুনির্বাচনি অধীক্ষার উত্তরপত্রে প্রশ্নের ত্রুটিক নম্বরের বিপরীতে প্রদত্ত বর্ণসম্বলিত বৃত্তসমূহ হতে সঠিক/সর্বোৎকৃষ্ট উত্তরের বৃত্তটি বলা পরেই কলাম দ্বারা সম্পূর্ণ করতে কর। প্রতিটি প্রশ্নের মান ১।]

প্রশ্নপত্রে কোনো প্রকার দাণ্ড/চিহ্ন দেয়া যাবে না।

- ১। চিকিৎসা বিজ্ঞানে ডিজিগনোজিসে কোনটি ব্যবহার করা হয়?

(ক) X-ray
(খ) IR
(গ) MRI
(ঘ) UV

উল্লিখিত থেকে ২ এবং ৩ নং প্রশ্নের উত্তর দাও :

মৌল	বহির্গতরের ইলেকট্রন বিন্যাস	সকল থেকে n এর মান 2
A	$(n-1)s^1$	
B	ns^2np^4	
C	$(n+1)s^2(n+1)p^4$	

A, B এবং C প্রদত্ত মৌলের প্রতীক নয়।

- ২। A_2B ও A_2C যৌত অবস্থার ত্রিমাত্রিক ক্রয়—

(ক) আয়নিক বন্ধন
(খ) সমযোজী বন্ধন
(গ) হাইড্রোজেন বন্ধন
(ঘ) সলিড বন্ধন

- ৩। A_2B যৌলের ক্ষেত্রে—

I. sp^3 সংকরণ ঘটে
II. বন্ধন কোণ 104.5°
III. $(L_p - L_p)$ বিকর্ষণ বিদ্যমান

নিচের কোনটি সঠিক?

(ক) I ও II
(খ) II ও III
(গ) I ও III
(ঘ) I, II ও III

- ৪। নিচের মৌলগুলোর ১ম আয়নীকরণ বিভবের সঠিক ক্রম কোনটি?

(ক) $Be > B > N > O$
(খ) $N > O > Be > B$
(গ) $O > N > B > Be$
(ঘ) $B > Be > N > O$

- ৫। হেলিনের প্রকৃত ভর কত?

(ক) $1.60 \times 10^{-24} g$
(খ) $1.66 \times 10^{-24} g$
(গ) $1.673 \times 10^{-24} g$
(ঘ) $1.675 \times 10^{-24} g$

- ৬। কোনটি প্রথম অক্সাইড?

(ক) NO
(খ) N_2O_5
(গ) N_2O_3
(ঘ) NO_2

- ৭। S-গ্রক মৌলের সংখ্যা কতটি?

(ক) 7
(খ) 12
(গ) 14
(ঘ) 16

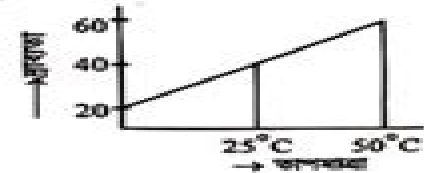
- ৮। পরমাণুর ৩য় শক্তিস্তরে ঘূর্ণিত ইলেকট্রনের কৌণিক ভরবেগ এর মান কোনটি?

(ক) $\frac{nh}{2\pi}$
(খ) $\frac{3h}{2\pi}$
(গ) $\frac{2h}{4\pi}$
(ঘ) $\frac{9h}{2\pi}$

- ৯। কোনটি উভমূলকী অক্সাইড?

(ক) Al_2O_3
(খ) Na_2O_2
(গ) KO_2
(ঘ) Fe_2O_3

উল্লিখিত হতে ১০ এবং ১১ নং প্রশ্নের উত্তর দাও :



- ১০। 200g NH_4Cl এর সম্পূর্ণ দ্রবক 50°C থেকে 25°C তাপমাত্রায় ঠাণ্ডা করলে কত গ্রাম NH_4Cl ক্রিস্টালিট হবে?

(ক) 25
(খ) 50
(গ) 75
(ঘ) 90

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Arshad Iqbal



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molar volume molecular ions molecular solids molecules moles Moseley law neutron properties non ideal behavior of gases orbital concept partial pressure calculations phase changes energies photons wave number Planck s quantum theory plasma state positive and negative ions pressure units properties of cathode rays covalent crystals properties of crystalline solids properties of positive rays quantum numbers quantum theory relative abundance Rutherford model of atom shapes of orbitals solid iodine structure solids properties solvent extraction spectrometer spin quantum number states of matter stoichiometry sublimation thermometry scales types of solids unit cell van der Waals equation vapor pressure and spectrum

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