OCTOBER - 2014

[Marks: 70

Time: 3 hrs.l

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question:

[7]

- (i) The Arrhenius equation is (a) $k = A e^{E_a}$ (d) $k = A e^{\frac{-E_s}{RT}}$ (c) $k = A e^{-E_a}$
- (ii) If the enthalpy of vaporisation of water at 100oC is 186.5 J. mol-1, the entropy of vaporization will be (b) 3.0 J . K⁻¹, mol⁻¹

(a) 4.0 J . K⁻¹, mol⁻¹ (c) 1.5 J , K-1, mol-1

(d) 0.5 J . K⁻¹, mol⁻¹

(iii) The atomicity of sulphur in orthorhombic sulphur is ___ (b) 6

(a) 8:

(d) 2

(iv) The major binding force in diamond is ___

(a) Covalent bond (c) Metallic bond

(b) Ionic bond (d) Co-ordinate covalent bond

(v) The boiling point of water at high altitude is low, because

(a) the temperature is low.

(b) the atmospheric pressure is low.

(c) the temperature is high.

- (d) the atmospheric pressure is high.
- (vi) The molar conductivity of cation and anion of salt BA are 180 and 220 mhos respectively. The molar conductivity of salt BA at infinite dilution is __ (a) 90 mhos.cm². (b) 110 mhos.cm².mol⁻¹

(c) 200 mhos.cm².mol⁻¹

- (d) 400 mhos.cm².mol⁻¹
- (vii) What is the process in which concentrated ore is reduced to the corresponding metal by heating at high temperature with a reducing agent? (a) Polling (b) Pyrometallurgy (c) Hydrometallurgy (d) Calcination
- Q. 2. Answer any THREE of the following:

[9]

(i) Describe anomalous behaviour of oxygen as compared with other elements of group 16 with reference to -

(a) Magnetic property

(b) Oxidation state (c) Hydrides

- (ii) What is the value of ΔS_{surr} for the following reaction at 298 K $6CO_{2(g)} + 6H_2O_{(f)} \longrightarrow C_6H_{12}O_{6(g)} + 6O_{2(g)}$ Given that : $\Delta G^{\circ} = 2879 \text{ kJ mol}^{-1}$, $\Delta S = -210 \text{ JK}^{-1} \text{ mol}^{-1}$.
- (iii) Sucrose decomposes in acid solution to give glucose and fructose according to the first order rate law. The half life of the reaction is 3 hours. Calculate fraction of sucrose which will remain after 8 hours.
- (iv) A solution containing 0.73 g of camphor (molar mass 152 g . mol⁻¹) in 36.8 g of acetone (boiling point 56.3°C) boils at 56.55°C. A solution of 0.564 g of unknown compound in the same weight of acetone boils at 56.46oC. Calculate the molar mass of the unknown compound.

Hsc 2014 Chemistry Paper

Maureen Slade, Roger Slade

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Cambridge Checkpoints HSC Chemistry 2017-19 Maureen Slade, Roger Slade, 2016-06-29 Material Science, Civil Engineering and Architecture Science, Mechanical Engineering and Manufacturing Technology II H.W. Liu, G. Wang, G.W. Zhang, 2014-09-30 Selected peer reviewed papers from the 2014 3rd International Conference on Advanced Engineering Materials and Architecture Science ICAEMAS 2014 July 26 27 2014 Huhhot Inner Mongolia China Energy Technology 2014 Cong Wang, Jan de Bakker, Cynthia K. Belt, Animesh Jha, Neale R. Neelameggham, Soobhankar Pati, Leon H. Prentice, Gabriella Tranell, Kyle S. Brinkman, 2013-12-30 This book addresses the issues intricacies and challenges relating to energy and environmental sciences Papers cover various technological aspects of sustainable energy eco systems and processes that improve energy efficiency and reduce and sequestrate carbon dioxide CO2 and other greenhouse emissions The collection also emphasizes the need for sustainable technologies in extractive metallurgy materials processing and manufacturing industries with reduced energy consumption and CO2 emission Industrial energy efficient technologies include innovative ore beneficiation smelting technologies recycling and waste heat recovery The book also carries contributions from all areas of non nuclear and non traditional energy sources including renewable energy sources such as solar wind and biomass Papers from the following symposia are presented in the book Energy Technologies and Carbon DioxideManagement High temperature Material Systems for Energy Conversion and Storage Solar Cell Silicon

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