OCTOBER - 2014

[Marks: 70

Time: 3 hrs.l

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question:

[7]

- (i) The Arrhenius equation is (a) $k = A e^{E_a}$ (d) $k = A e^{\frac{-E_s}{RT}}$ (c) $k = A e^{-E_a}$
- (ii) If the enthalpy of vaporisation of water at 100oC is 186.5 J. mol-1, the entropy of vaporization will be (b) 3.0 J . K⁻¹, mol⁻¹

(a) 4.0 J . K⁻¹, mol⁻¹ (c) 1.5 J , K-1, mol-1

(d) 0.5 J . K⁻¹, mol⁻¹

(iii) The atomicity of sulphur in orthorhombic sulphur is ___ (b) 6

(a) 8:

(d) 2

(iv) The major binding force in diamond is ___

(a) Covalent bond (c) Metallic bond

(b) Ionic bond (d) Co-ordinate covalent bond

(v) The boiling point of water at high altitude is low, because

(a) the temperature is low.

(b) the atmospheric pressure is low.

(c) the temperature is high.

- (d) the atmospheric pressure is high.
- (vi) The molar conductivity of cation and anion of salt BA are 180 and 220 mhos respectively. The molar conductivity of salt BA at infinite dilution is __ (a) 90 mhos.cm². (b) 110 mhos.cm².mol⁻¹

(c) 200 mhos.cm².mol⁻¹

- (d) 400 mhos.cm².mol⁻¹
- (vii) What is the process in which concentrated ore is reduced to the corresponding metal by heating at high temperature with a reducing agent? (a) Polling (b) Pyrometallurgy (c) Hydrometallurgy (d) Calcination
- Q. 2. Answer any THREE of the following:

[9]

(i) Describe anomalous behaviour of oxygen as compared with other elements of group 16 with reference to -

(a) Magnetic property

(b) Oxidation state (c) Hydrides

- (ii) What is the value of ΔS_{surr} for the following reaction at 298 K $6CO_{2(g)} + 6H_2O_{(f)} \longrightarrow C_6H_{12}O_{6(g)} + 6O_{2(g)}$ Given that : $\Delta G^{\circ} = 2879 \text{ kJ mol}^{-1}$, $\Delta S = -210 \text{ JK}^{-1} \text{ mol}^{-1}$.
- (iii) Sucrose decomposes in acid solution to give glucose and fructose according to the first order rate law. The half life of the reaction is 3 hours. Calculate fraction of sucrose which will remain after 8 hours.
- (iv) A solution containing 0.73 g of camphor (molar mass 152 g . mol⁻¹) in 36.8 g of acetone (boiling point 56.3°C) boils at 56.55°C. A solution of 0.564 g of unknown compound in the same weight of acetone boils at 56.46oC. Calculate the molar mass of the unknown compound.

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- 1. Understanding the eBook Hsc October 2014 Chemistry Question Paper
 - The Rise of Digital Reading Hsc October 2014 Chemistry Question Paper
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Hsc October 2014 Chemistry Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Hsc October 2014 Chemistry Question Paper
 - User-Friendly Interface

- 4. Exploring eBook Recommendations from Hsc October 2014 Chemistry Question Paper
 - Personalized Recommendations
 - Hsc October 2014 Chemistry Question Paper User Reviews and Ratings
 - Hsc October 2014 Chemistry Question Paper and Bestseller Lists
- 5. Accessing Hsc October 2014 Chemistry Question Paper Free and Paid eBooks
 - Hsc October 2014 Chemistry Question Paper Public Domain eBooks
 - Hsc October 2014 Chemistry Question Paper eBook Subscription Services
 - Hsc October 2014 Chemistry Question Paper Budget-Friendly Options
- 6. Navigating Hsc October 2014 Chemistry Question Paper eBook Formats
 - o ePub, PDF, MOBI, and More
 - Hsc October 2014 Chemistry Question Paper Compatibility with Devices
 - Hsc October 2014 Chemistry Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Hsc October 2014 Chemistry Question Paper
 - Highlighting and Note-Taking Hsc October 2014 Chemistry Question Paper
 - Interactive Elements Hsc October 2014 Chemistry Question Paper
- 8. Staying Engaged with Hsc October 2014 Chemistry Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Hsc October 2014 Chemistry Question Paper
- 9. Balancing eBooks and Physical Books Hsc October 2014 Chemistry Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Hsc October 2014 Chemistry Question Paper
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Hsc October 2014 Chemistry Question Paper
 - Setting Reading Goals Hsc October 2014 Chemistry Question Paper
 - Carving Out Dedicated Reading Time

- 12. Sourcing Reliable Information of Hsc October 2014 Chemistry Question Paper
 - Fact-Checking eBook Content of Hsc October 2014 Chemistry Question Paper
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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