

STD XII - CHEMISTRY

Time : 3 hrs.]

MARCH - 2013

[Marks : 70

SECTION - I

Q. 1. Select and write the most appropriate answers from the given alternatives for each sub-question :

[7]

- (i) In body centred cubic structure the space occupied is about _____.
 (a) 68% (b) 53% (c) 38% (d) 32%
- (ii) For a gaseous reaction the unit of rate of reaction is _____.
 (a) L atm s^{-1} (b) $\text{atm mol}^{-1} \text{s}^{-1}$ (c) atm s^{-1} (d) mol s
- (iii) Which of the following compounds contains S = O as well as S = S bonds?
 (a) Sulphuric acid (b) Thiosulphuric acid
 (c) Sulphurous acid (d) Thiosulphurous acid
- (iv) Which of the following solutions shows maximum depression in freezing point?
 (a) 0.5 M Li_2SO_4 (b) 1 M NaCl (c) 0.5 M $\text{Al}_2(\text{SO}_4)_3$ (d) 0.5 M BaCl_2
- (v) For a chemical reaction $dS = -0.035 \text{ kJ/K}$ and $dH = -20 \text{ kJ}$. At what temperature does the reaction turn nonspontaneous?
 (a) 5.14 K (b) 57.14 K (c) 571.4 K (d) 5714.0 K
- (vi) The standard e.m.f. of the following cell is 0.463 V

$$\text{Cu}_{(s)} \mid \text{Cu}_{(aq)}^{++} (1\text{M}) \parallel \text{Ag}_{(aq)}^{++} (1\text{M}) \mid \text{Ag}_{(s)}$$
 . What is the standard potential of Cu electrode?
 (a) 1.137 V (b) 0.337 V (c) 0.463 V (d) -0.463 V
- (vii) Fe_2O_3 is reduced to spongy iron near the top of blast furnace by _____.
 (a) H_2 (b) CaO (c) SiO_2 (d) CO

Q. 2. Answer any SIX of the following :

[12]

- (i) Distinguish between crystalline solid and amorphous solid.
- (ii) State Kohlrausch Law and write mathematical expression of molar conductivity of the given solution at infinite dilution.
- (iii) Write cell reactions in lead storage battery during discharge.
- (iv) Draw structures and write geometry of PCl_3 and PCl_5 .
- (v) Prove that $\Delta H = \Delta U + \Delta nRT$. What is the condition under which $\Delta U = \Delta H$?
- (vi) Mention names and formulae of two ores of aluminium.
- (vii) Derive the relationship between relative lowering of vapour pressure and molar mass of nonvolatile solute.
- (viii) What is pseudo first order reaction? Give one example of it.

Q. 3. Answer any THREE of the following :

[9]

- (i) Calculate the mole fraction and molality of HNO_3 in a solution containing 12.2% HNO_3 . (Given - atomic masses : H = 1, N = 14, O = 16)
- (ii) Consider the reaction -



What are the values of the following at the same time ?

- (a) $-\frac{d[\text{I}^{-}]}{dt}$ (b) $-\frac{d[\text{S}_2\text{O}_8^{2-}]}{dt}$ (c) $-\frac{d[\text{I}_3^{-}]}{dt}$

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acid base ph titration introduction the royal society of chemistry - May 02 2023

web ph titration graph 1 lab data provide columns a and b 2 a graph of ph column b plotted as ordinate vs volume column a as abscissa provides a titration curve as illustrated by graph 1 or 2 below 3 estimate of the midpoint of the steeply rising section of the titration curve provides an evaluation of the equivalence point

acid base titration lab report 260678496 pdf studocu - Feb 16 2022

web acid base titration lab report jessica chung chem 267 lab date completed 25 september 2019 date turned in 9 october 2019 figure 1 acid base titration curve from titration with indicator for part 1 volume of naoh added ml ph of the solution 4 5 4 6 4 6 5 9 5 9

experiment 3 acid base titration chem 1202 expt 3 title - Sep 25 2022

web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

acid base titration curves 1 1 microlab inc - Apr 01 2023

web background plot of the ph of a solution against the volume of titrant added is called a titration curve the ph can be measured directly with a ph meter while titrant is added from a buret for the acids used in this experiment the titrant will always be a 0.1 M solution of the strong base naoh

17 3 acid base titrations chemistry libretexts - Oct 27 2022

web oct 30 2022 there are two basic types of acid base titrations indicator and potentiometric in an indicator based titration you add another chemical that changes color at the ph equal to the equivalence point when the acid and base are in

titration screen experiment resource rsc education - Aug 25 2022

web rp 1 make up a volumetric solution and carry out a simple acid base titration physical chemistry acids and bases ph

curves titrations and indicators titrations of acids with bases students should be able to perform calculations for these titrations based on experimental results inorganic chemistry transition metals variable

acid base titration chemistry 1210 lab report containing an studocu - Sep 06 2023

web strong acids and strong bases completely ionize in solution resulting in water and a salt for a weak acid there s only partial ionization the titration curve of a strong base strong acid showed a stretched out curve as it started with a slow gradual change in ph as it reached the equivalence point