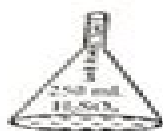


অধ্যায় ১: ল্যাবরেটরির নিরাপদ ব্যবহার

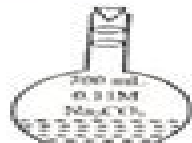


পটাসিয়াম কপ্টার, এখানে মোট পটাসিয়াম, আরও কপ্টারের শীর্ষস্থানীয় কপ্টারের পটাসিয়াম এবং মনোমের টেম্পের পটাসিয়াম
পটাসিয়াম উত্তর অধ্যায়ভিত্তিক মেট্রিক হলে। এগুলোর অনুশীলন করলে তুমি এ অধ্যায়টি থেকে সেরেফেরে পটাসিয়াম কপ্টারের
এদের উত্তর লিখতে পারবে সহজেই।

প্রশ্ন ১



পাত্র-A



পাত্র-B

সি. কে. ১০।এস. ১।

- মোলের প্রকণ কী?
- "কেলস পটাসিয়াম লবণের প্রকণতা, 'অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়'— ব্যাখ্যা করো।
- A-পাত্রের বিন্যাসন এলিটের ভর কত নির্ণয় করো।
- প্রকণ দুটি মিশ্রিত করলে, মিশ্রণটি কী পরিমাণ অক্সিজেন দ্বারা পরিণত হবে— বিশ্লেষণ করো।

১ নং প্রশ্নের উত্তর

1000 mL প্রকণে কোন প্রকণের 1 mol প্রকণিত থাকলে উক্ত প্রকণকে মোলার প্রকণ বলে।

কেলস পটাসিয়াম লবণের প্রকণতা অপমাত্রা বৃদ্ধির সাথে সাথে এক পর্যায়ে বৃদ্ধি না পেয়ে হ্রাস পায়। এর কারণ, তাপমাত্রা বৃদ্ধির সাথে সাথে প্রকণতা বৃদ্ধি পেতে থাকলেও একটি নির্দিষ্ট তাপমাত্রার কেলস পটাসিয়াম লবণ নিষ্কৃত হয়। তখন প্রকণতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর কথা বলা যায়। $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ এর ক্ষেত্রে পটাসিয়ামে প্রকণিত করার সময় বিশ্লেষণ প্রক্রিয়াটি তাপমাত্রা 32.4°C তাপমাত্রার উপরে $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ যখন Na_2SO_4 এ পরিণত হয় তখন বিশ্লেষণ প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে 32.4°C তাপমাত্রার উপরে তাপমাত্রা বৃদ্ধিজনিত কারণে প্রকণতার হ্রাস ঘটে।

২

আমরা জানি,
 $w = SMV$
 $= (0.5 \times 98 \times 0.25) \text{ g}$
 $= 12.25 \text{ g}$

অতএব, প্রকণিত H_2SO_4 এর ভর
 $= 12.25 \text{ g}$

উদাহরণের A-পাত্রের অবস্থিতি,
 H_2SO_4 এর আণবিক ভর,
 $M = 1 \times 2 + 32 + 16 \times 4$
 $= 2 + 32 + 64$
 $= 98 \text{ g/mol}$

প্রকণের ঘনমাত্রা,
 $S = 0.5 \text{ M}$
প্রকণের আয়তন,
 $V = 250 \text{ mL}$
 $= \frac{250}{1000} = 0.25 \text{ L}$

প্রকণিত H_2SO_4 এর ভর, $w = ?$

প্রশ্ন-A: H_2SO_4 এর মোল সংখ্যা $= \left(0.5 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.125 \text{ mol}$

প্রশ্ন-B: Na_2CO_3 এর মোল সংখ্যা $= \left(0.11 \times \frac{250}{1000}\right) \text{ mol}$
 $= 0.077 \text{ mol}$

সম্প্রতি বিক্রিয়া: $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$

সমীকরণ হতে পাই, 1 mol $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol Na_2CO_3 বিক্রিয়া করে H_2SO_4 এর 1 mol এর সাথে
 0.077 mol

$(0.077) \text{ mol} = 0.077 \text{ mol}$

অবশিষ্ট H_2SO_4 এর পরিমাণ $= (0.125 - 0.077) \text{ mol}$
 $= 0.048 \text{ mol}$

প্রশ্ন-A এর প্রকণের সাথে প্রশ্ন-B এর প্রকণের বিক্রিয়ার ফলে 0.048 mol H_2SO_4 অবশিষ্ট থাকায় মিশ্রিত প্রকণের প্রকৃতি হবে অক্সিজেন।

প্রশ্ন ২ একদল ছাত্র ল্যাবরেটরিতে X-পদ্ধতিতে 10 mg এর প্রকণ প্রকণিত এবং Y-পদ্ধতিতে 0.5 g পরিমাণ 2% $\frac{W}{V}$ HCl এর প্রকণ প্রকণিত করে।

সি. কে. ১০।এস. ১।

- নোড কাকে বলে?
- Cl_2 অপোনার— ব্যাখ্যা করো।
- উদাহরণের এলিটটির মোলার ঘনমাত্রা নির্ণয় করো।
- উদাহরণের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

২ নং প্রশ্নের উত্তর

নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেক্ট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সমান হবে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হবে সমযোজী বন্ধন। অপরদিকে, আয়নিক বন্ধনে আবদ্ধ পরমাণুর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর বেশি হবে অপরদিকে হয়। যেহেতু Cl_2 এর তড়িৎ ঋণাত্মকতা 3.0 ও O এর তড়িৎ ঋণাত্মকতা 3.5 তাই Cl_2 বন্ধন সমযোজী ও অপরদিকে যৌগ। উল্লেখ্য যৌগিক অণুর ক্ষেত্রে এটি প্রযোজ্য থাকে।

উদাহরণে বর্ণিত এলিটটি হলো 2% $\left(\frac{W}{V}\right) \text{ HCl}$

2% $\left(\frac{W}{V}\right) \text{ HCl}$ এর অর্থ হলো—

100 mL প্রকণে HCl এর ভর 2g

∴ 1000 mL " " " " $\left(\frac{2 \times 1000}{100}\right) \text{ g}$
 $= 20 \text{ g}$

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