

# On your own

1. One gallon of gasoline can be considered to contain about 25 moles of octane ( $C_8H_{18}$ ). Octane burns according to the following equation:



- (a) How many moles of oxygen gas must be used to burn a gallon of gasoline?  $312.5 \text{ mol } O_2$   
 (b) How many moles of carbon dioxide are formed?  $200 \text{ mol } CO_2$   
 (c) How many grams of carbon dioxide are formed?  $8500 \text{ g } CO_2$

2. Many statues are made of marble (calcium carbonate). Acid rain contains sulfuric acid, which reacts with marble according to the following equation:



- (a) How many moles of marble will be consumed if 5 mol  $H_2SO_4$  fall onto the statue?  $5 \text{ mol } CaCO_3$   
~~How many moles of  $CO_2$  are produced when 12 g marble are reacted at STP?~~  
 (c) How many grams of the statue will be consumed if 500 g  $H_2SO_4$  fall onto the statue?  $510 \text{ g } CaCO_3$

3. Plants convert carbon dioxide, water, and energy from the sun into sugar ( $C_6H_{12}O_6$ ) and dioxygen in the process of photosynthesis.



- (a) How many grams of sugar are produced from 1000g (one liter) of water?  $1667 \text{ g } C_6H_{12}O_6$   
~~How many liters of  $CO_2$  are needed to react with one liter of water?~~

4. Thermite reactions have been used in the past for welding railroad rails, making incendiary bombs, and igniting solid-fuel rocket motors.



- (a) How many grams of iron (III) oxide will be needed to react in order to form 15.0 g iron?  $21.4 \text{ g } Fe$   
 (b) How many grams of aluminum will be needed to react in order to form 15.0 g iron?  $7.2 \text{ g } Al$   
 (c) How many grams of aluminum oxide will be formed if  $1.63 \times 10^{23}$  atoms iron are formed?  $13.8 \text{ g } Al_2O_3$

5. Calcium cyanamide ( $CaCN_2$ ) is used as a fertilizer. The water in soil react with the fertilizer to deposit  $NH_3$  and  $CaCO_3$  into the ground.



- (a) How many grams of  $NH_3$  are deposited into the soil if 250 kg  $CaCN_2$  are used?  $106250 \text{ g } NH_3$   
 (b) How many grams of water from the soil must react with 250 kg  $CaCN_2$ ?  $118750 \text{ g } H_2O$

# Hs Chemistry Pogil Activity Basic Stoichiometry On Your Own

**Jennifer Louise Pakkala**



## **Hs Chemistry Pogil Activity Basic Stoichiometry On Your Own:**

*Calculating in Chemistry* Cosmas C Mbuji, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities. Students in higher courses may find it a helpful revision and enhance their clarity. The book completely discusses the key topics in basic stoichiometry including the mole concept, reacting quantities, and empirical and molecular formulas. It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance. This is particularly useful to a beginner. Chapter 2 describes how to calculate reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems. The chapter also describes and applies the concept of a limiting reagent. Two methods of determining a limiting reagent are explained and illustrated. The concept of limiting reagent is extended to reacting volumes of gases with a short cut method. A short cut method for solving reacting quantities involving masses and volumes of gases is also given. Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas. A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner. In chapter 4 we discuss percentage yield and purity. A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances. A student will find these examples helpful in relating different formulas for percentage purity. The last chapter introduces a graphical method for reacting quantities. The method may provide a new way of looking at chemical reactions. Examples are given to illustrate the method including how it can be used to determine limiting reagents. It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course. Teachers may also find this book a good resource for their lessons in stoichiometry. **STOICHIOMETRY NARAYAN**

CHANGDER, 2024-04-01 Note: Anyone can request the PDF version of this practice set workbook by emailing me at cbsenet4u@gmail.com. You can also get full PDF books in quiz format on our youtube channel <https://www.youtube.com/smartquiziz>. I will send you a PDF version of this workbook. This book has been designed for candidates preparing for various competitive examinations. It contains many objective questions specifically designed for different exams. Answer keys are provided at the end of each page. It will undoubtedly serve as the best preparation material for aspirants. This book is an engaging quiz eBook for all and offers something for everyone. This book will satisfy the curiosity of most students while also challenging their trivia skills and introducing them to new information. Use this invaluable book to test your subject matter expertise. Multiple choice exams are a common assessment method that all prospective candidates must be familiar with in today's academic environment. Although the majority of students are accustomed to this MCQ format, many are not well versed in it. To achieve success in MCQ tests, quizzes, and trivia challenges, one requires test taking techniques and skills in addition to subject knowledge. It also provides you with the skills and information you need to achieve a good score in challenging tests or competitive examinations. Whether you have studied the subject on your own, read for pleasure, or completed coursework, it

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*Supporting Traditional Methods with a Constructivist Approach to Learning* Kenneth Munoz Abayan, ProQuest Dissertations and Theses (Electronic resource collection), 2014 Stoichiometry is a fundamental topic in chemistry that measures a quantifiable relationship between atoms molecules etc Stoichiometry is usually taught using expository teaching methods Students are passively given information in the hopes they will retain the transmission of information to be able to solve stoichiometry problems masterfully Cognitive science research has shown that this kind of instructional teaching method is not very effecting in meaningful learning practice Instead students must take ownership of their learning The students need to actively construct their own knowledge by receiving interpreting integrating and reorganizing that information into their own mental schemas In the absence of active learning practices tools must be created in such a way to be able to scaffold difficult problems by encoding opportunities necessary to make the construction of knowledge memorable thereby creating a usable knowledge base Using an online e learning tool and its potential to create a dynamic and interactive learning environment may facilitate the learning of stoichiometry The study entailed requests from volunteer students IRB consent form a baseline questionnaire random assignment of treatment pre and post test assessment and post assessment survey These activities were given online A stoichiometry based assessment was given in a proctored examination at the University of Texas at Arlington UTA campus The volunteer students who took part in these studies were at least 18 of age and were enrolled in General Chemistry 1441 at the University of Texas at Arlington Each participant gave their informed consent to use their data in the following study Students were randomly assigned to one of 4 treatment groups based on teaching methodology Dimensional Analysis Operational Method Ratios and Proportions and a control group who just received instruction through lecture only In this study an e learning tool was created to demonstrate several methodologies on how to solve stoichiometry which are all supported by chemical education research Comparisons of student performance based on pre and post test assessment and a stoichiometry based examination was done to determine if the information provided within the e learning tool yielded greater learning outcomes compared to the students in the absence of scaffold learning material The e learning tool was created to help scaffold the problem solving process necessary to help students N 394 solve stoichiometry problems Therein the study investigated possible predictors for success on a stoichiometry based examination students conceptual understanding of solving stoichiometry problems and their explanation of reasoning It was found that the way the student answered a given stoichiometry question i e whether the student used

dimensional analysis operational method or any other process was not statistically relevant  $p = 0.05$ . More importantly if the students were able to describe their thought process clearly these students scored significantly higher on stoichiometry test mean 84  $p$

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