

Chapter 1: Stoichiometry

Q4.

Sol: $\text{CaCO}_3 \longrightarrow \text{CaO} + \text{CO}_2$

Since 1 mole of CaCO_3 produces 1 mole of CaO , So

1 mole of CaO will produce 1 mole of CaCO_3 .

\therefore 1 mole of CaO \longrightarrow 1 mole of CaCO_3 .

Sol: $\text{BaCl}_2 \longrightarrow \text{BaCO}_3$

Given:

5 moles of BaCl_2 contains 5 moles of Ba & 10 moles of Cl & each form 5 moles of BaCO_3 , \therefore each 1 mole of BaCO_3 contains 1 mole of Ba .

Sol: $2 \text{Na} \longrightarrow \text{Na}_2\text{O}$

5 moles O_2 (Given)

\therefore 2 moles produces 1 mole of Na_2O

$\therefore 5 \longrightarrow \frac{5}{2} \times 2 = 5$

$\therefore 5 \longrightarrow 5 \times \frac{1}{2} = 2.5$ moles of Na_2O .

Q1.

Ans: $\text{Ca} + \frac{1}{2} \text{O}_2 \longrightarrow \text{CaO}$ (balanced chemical equation)

Wt. of $\text{Ca} = 40 \text{ g/mole}$

So, moles of $\text{Ca} = \frac{80}{40} = 2 \text{ moles}$

From above balanced chemical equation, 2 moles of Ca react with $\frac{1}{2} \times 2 = 1$ moles of O_2 .

Wt. of O_2 with which 80g of Ca react $= \frac{1}{2} \times 32 = 16 \text{ g}$

Ans.

Q2.

Sol: $\text{NaCl} \longrightarrow \text{NaCl}$

58 gm

If we have not balanced the chemical equation, then we have to apply PMOL rule as per equivalence ratio.

By PMOL rule Cl atom

From NaCl atom to be reacted \therefore 58 gm of Cl (molecular weight)

Wt. of moles of $\text{NaCl} = \text{Wt.} = 2 \times \text{wt. of moles of } \text{Cl}_2 = 71$

$\frac{58}{71} = 0.817$ moles of Cl_2

moles of $\text{Cl}_2 = \frac{58}{71}$

Modern Approach To Chemical Calculations Solutions

Albert A Gayle



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