

Chapter 2: The Chemistry of life

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Part 1: Atoms make up molecules

- All things made up of atoms
 - All matter that takes up space (solid, liquids, gases) is made up of elements
 - Element: Substance that cannot be broken down chemically into other substances
 - EX: Oxygen, gold, aluminum, calcium, iron
 - Over 100 elements we know of (only 25 essential to life)
 - Smallest unit of an element = an **atom**
- Atoms themselves are all made up of 3 components:
 - **Protons** = positive charge (+) (located in atomic nucleus)
 - **Neutrons** = no charge (located in atomic nucleus)
 - **Electrons** = negative charge (-)
- Atoms electrically neutral b/c the number of protons is equal to number of electrons
- **Protons:**
 - Each element has a unique number of protons
 - EX: Hydrogen ALWAYS has one
 - EX: Sodium ALWAYS has two
 - **Atomic number** = the number of protons an atom has
 - **Atomic mass** = number of protons + neutrons
 - Electrons so light we ignore them in biology
- **Neutrons:**
 - Number of protons never changes but number of neutrons an atom has can change
 - **Isotope** = an atom with a different number of neutrons (not the same number)
 - Radiation comes from isotope which some are stable and some are unstable
- **Radioactivity/radiation benefits:**
 - Harvested for energy
 - Treatment for cancer
 - Age of fossils
 - Medical imaging
- **Electrons:**
 - Distributed around the atomic nucleus in shells (constant motion)
 - First electron shell holds up to 2 electrons
 - Second electron shell holds up to 8 electrons
 - Atoms are most stable when the outer shell of the atom is full
 - To fill their shells atoms, lose, gain, or share electrons
- Atoms that lose or gain electrons become stable are called **ions** (electrons don't match).
 - Atom that loses an electron will have a positive charge = **cation**
 - Atom that gains an electron will have a negative charge = **anion**
- Atoms link together using chemical bonds to form molecules

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