

Pre-Ap Chemistry Unit 6 – review for finals

These are the topics you should know and be able to answer questions about:

1. Chemical reactions

- What are some examples of evidence of chemical reactions?
 - Evidence that a chemical change *may have occurred* can include:
 - Production of bubbles (gas)
 - Formation of a precipitate (solid)
 - Production of heat and light
 - Color change
 - Noise
 - Ultimately, only the formation of a **NEW** substance can be absolute proof that a chemical change has occurred.
- What is the difference between chemical and physical reactions?
 - Chemical reactions produce new products – what you start with and end with are different
 - Physical reactions change the appearance of the substance, but you still have the same substance that you start with

2. Balance equations.

- Be able to balance equations
 - Either use math or tiddlywinks or draw a picture
- Why do we have to balance equations
 - Due to the law of conservation of mass. Mass is not created or destroyed during a chemical reaction. Due to this, the mass of the reactants must equal the mass of the products. Which means that there must be a balanced mass on each side of the reaction
- What are we trying to balance when we balance equations?
 - We need the same number of atoms of each element on each side of the arrow
- If you have balanced an equation and notice that the coefficients can be reduced, what do you? And why?
 - Reduce the coefficients.
 - We need the lowest ratio of coefficients
- Draw for me the particle representation of a balanced equation (circles) for the formation of water ($\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$)



- In order to balance the equation $\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$ the correct coefficient for oxygen would be:
— i. 3

3. Types of reactions

- What are the different types of reactions?
 - Synthesis
 - Decomposition
 - Single replacement
 - Double replacement
 - Combustion

Modeling Chemistry Unit 6 Review V1

Sebastian Brünink



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