

Chemistry – Unit 5 Worksheet 3 Empirical and Molecular Formulas

Show all your work, show all steps, use following guidelines. Do not forget to show units and label your answers.

1. Find the empirical formula of a compound containing 40.0 g of bromine and 4.0 g of hydrogen.

$$\text{mole Br} = \frac{40.0 \text{ g}}{79.90 \text{ g/mol}} = 0.501 \text{ mole Br}$$

1:1 ratio

$$\text{mole H} = \frac{4.0 \text{ g}}{1.008 \text{ g/mol}} = 3.97 \text{ mole H}$$



2. What is the empirical formula of a hydrocarbon compound if 10.0 g of carbon and 0.75 g of hydrogen are present?

$$\text{mole C} = \frac{10.0 \text{ g}}{12.01 \text{ g/mol}} = 0.833 \text{ mole C}$$

$$\text{mole H} = \frac{0.75 \text{ g}}{1.008 \text{ g/mol}} = 0.744 \text{ mole H}$$

1:1 ratio

$$\text{CH}_1 = \frac{0.744}{0.833} = 0.893 \text{ mole H}$$



3. A compound was identified and found to contain 40.0 g of hydrogen, 80.0 g of hydrogen, and 80.0 g of oxygen. What is the empirical formula of this compound?

$$\text{mole H} = \frac{80.0 \text{ g}}{1.008 \text{ g/mol}} = 79.4 \text{ mole H}$$

$$\text{mole H} = \frac{80.0 \text{ g}}{1.008 \text{ g/mol}} = 79.4 \text{ mole H}$$

1:1:1 ratio

$$\text{mole O} = \frac{80.0 \text{ g}}{16.00 \text{ g/mol}} = 5.00 \text{ mole O}$$



4. A compound composed of hydrogen and oxygen is found to contain 10.0 g of hydrogen and 80.0 g of oxygen. What is the ratio of these compounds? Is it a gas? Find the empirical molecular formula.

$$\text{mole H} = \frac{10.0 \text{ g}}{1.008 \text{ g/mol}} = 9.92 \text{ mole H}$$

$$\text{mole O} = \frac{80.0 \text{ g}}{16.00 \text{ g/mol}} = 5.00 \text{ mole O}$$

1:1 = 10.0 g hydrogen : 80.0 g oxygen

10.0 g H : 80.0 g O = 1:8 ratio

10.0 g H : 80.0 g O = 1:8 ratio

10.0 g H : 80.0 g O = 1:8 ratio

10.0 g H : 80.0 g O = 1:8 ratio

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