

# Test Quantities in Chemical Reactions SCH3U #13

Name: \_\_\_\_\_

\_\_\_\_K \_\_\_\_T \_\_\_\_C \_\_\_\_A

Assessment of Learning

(Product)

10

20

8

7

Multiple Choice: Circle the most correct answer. (10K)

1. What is the molar mass of  $\text{AuCl}_3$ ?

A) 96 g

B) 130 g

C) 232.5 g

D) 303.6 g

2. Determine the mass in grams of 0.0488 mol cobalt.

A) 0.00852 g

B) 0.294 g

C) 2.88 g

D) 0.00830 g

3. The number of molecules of hydrogen peroxide found in 3.15 mol  $\text{H}_2\text{O}_2$  is

A) 3

B)  $1.90 \times 10^{21}$

C)  $3.80 \times 10^{21}$

D)  $7.60 \times 10^{21}$

4. The percentage of oxygen, by mass, in  $\text{Mg}_3(\text{PO}_4)_2$  is

A) 15.9%

B) 23.8%

C) 27.7%

D) 48.7%

E) None of the above

5. An unknown gas is found to consist of 10.48 g of nitrogen and 11.96 g of oxygen. What is the percentage composition of this gas?

A) 28.6% N, 71.4 % O

B) 48.7%N, 51.3% O

C) 53.2% N, 46.7% O

D) 74.5% N, 25.5% O

6. A compound contains 38.7% K, 13.9% N, and 47.4% O by mass. What is the empirical formula of the compound?

A)  $\text{KNO}_3$

B)  $\text{K}_2\text{N}_2\text{O}_3$

C)  $\text{KNO}_2$

D)  $\text{K}_2\text{N}_2\text{O}_2$

E)  $\text{K}_4\text{NO}_3$

7. If the molar mass of a hydrocarbon is 26.0 g/mol, and its empirical formula is  $\text{CH}$ , its molecular formula is

A)  $\text{CH}$

B)  $\text{C}_2\text{H}_2$

C)  $\text{C}_3\text{H}_4$

D)  $\text{C}_4\text{H}_6$

E)  $\text{C}_6\text{H}_6$

8. Consider the following reaction:  $\text{Ca}(\text{OH})_{2(aq)} + \text{Al}_2(\text{SO}_4)_{3(aq)} \rightarrow \text{CaSO}_{4(s)} + \text{Al}(\text{OH})_{3(s)}$

When balanced, the mole ratio for calcium hydroxide reacting with aluminum sulphate is

A) 1:1

B) 2:1

C) 3:3

D) 1:2

E) 3:1

9. The limiting reactant in a chemical reaction is the \_\_\_\_\_, that is \_\_\_\_\_.

A) Reactant, in excess

B) Product, left over

C) Reactant, used up

D) Product, used up

# Modeling Chemistry Mole Test Answer

**Rob Ritchie**



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