

## Unit 7 Worksheet 1: Molar-mass calculations

Four marks will be given for each question.

1. Write down balanced chemical equations

a) for the reaction of propanoic acid, acetic acid and acetic acid with calcium hydroxide

b) for the reaction of calcium hydroxide with calcium hydroxide

c) for the reaction of calcium hydroxide with calcium hydroxide

2. Hydrogen sulfide gas, which smells like rotten eggs, burns to give two products. Write the balanced chemical equation. What mass of hydrogen sulfide gas is needed to produce 100 g of each of the products?

Equation: \_\_\_\_\_

Masses: \_\_\_\_\_

Change: \_\_\_\_\_

Answer: \_\_\_\_\_

3. Propane,  $C_3H_8$ , burns in air to form carbon dioxide and water. If 100 g of propane is burned, how many grams of propane are burned?

Equation: \_\_\_\_\_

Masses: \_\_\_\_\_

Change: \_\_\_\_\_

Answer: \_\_\_\_\_

4. Acetic acid,  $CH_3COOH$ , has a molar mass of 60 g/mol. If 100 g of acetic acid is burned, how many grams of acetic acid are burned? If 100 g of acetic acid is burned, how many grams of acetic acid are burned?

Equation: \_\_\_\_\_

Masses: \_\_\_\_\_

Change: \_\_\_\_\_

Answer: \_\_\_\_\_

# Modeling Chemistry U8 Test V2

**Sabine Zange**



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