

Ions that form soluble compounds	Exceptions	Ions that form soluble compounds	Exceptions
Group 1 ions (Li^+ , Na^+ , K^+ , ...)	-	Carbonate (CO_3^{2-})	When combined with group 1 ions or ammonium (NH_4^+)
Ammonium (NH_4^+)	-	Chromate (CrO_4^{2-})	When combined with group 1 ions, Ca^{2+} , Mg^{2+} , or ammonium (NH_4^+)
Nitrate (NO_3^-)	-	Phosphate (PO_4^{3-})	When combined with group 1 ions or ammonium (NH_4^+)
Acetate ($\text{C}_2\text{H}_3\text{O}_2^-$ or CH_3COO^-)	-	Sulfide (S^{2-})	When combined with group 1 ions or ammonium (NH_4^+)
Hydrogen Carbonate (HCO_3^-)	-	Hydroxide (OH^-)	When combined with group 1 ions, Ca^{2+} , Ba^{2+} , Sr^{2+} or ammonium (NH_4^+)
Chlorate (ClO_3^-)	-	Sulfates (SO_4^{2-})	When combined with Ag^+ , Ca^{2+} , Sr^{2+} , Ba^{2+} , and Pb^{2+}
Perchlorate (ClO_4^-)	-	Halides (Cl^- , Br^- , I^-)	When combined with Ag^+ , Pb^{2+} , and Hg_2^{2+}

Solubility is important because it directly impacts the amount of radioisotope available for the reaction. It may affect distribution in the radiation source resulting in dosimetry issues. For example, in Benega et. al. [34] a phosphorus-32 radioactive source for spinal cancer treatment was developed by mixing the radioisotope with a catalyst solution. This solution is then added to an epoxy resin. Both solutions need to result in a homogenous product. This won't be achievable if the isotope doesn't properly mix with the catalyst solution (water based).

The solubility product constant, K_{sp} , is the equilibrium constant for a solid substance dissolving in an aqueous solution. It represents the level at which a solute will dissolve in solution. Equation 17 shows the correlation.

$$aA_{(s)} \leftrightarrow bB_{(aq)} + cC_{(aq)} \quad (15)$$

Model 1 Rules Of Solubility In Aqueous Solutions

Raffaella Di Napoli



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