

**CLASS SET**

**Mixed Stoichiometry Problems**

1. How many moles of  $\text{H}_2$  would be required to completely react with  $\text{O}_2$  to produce 5 moles of water? **5 mol  $\text{H}_2$**
2.  $\text{H}_2\text{SO}_4 + \text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O}$ 
  - a. Balance this equation
  - b. What mass of  $\text{H}_2\text{SO}_4$  would be required to react with 0.75 mol of  $\text{NaOH}$ ? **37g**
3. What mass of  $\text{NO}_2$  is formed when  $\text{NO}$  reacts with 384 g of  $\text{O}_2$ ? **1104g**
4. If 30 g of  $\text{H}_2\text{SO}_4$  reacts with aluminum hydroxide in a double replacement reaction, what mass of water is produced? **11.0g**
5. If 10 L of  $\text{CH}_4$  combusts with  $\text{O}_2$ , what volume of water will be produced? **20 L  $\text{H}_2\text{O}$**
6. Tin(II) fluoride,  $\text{SnF}_2$ , is used in some toothpastes. It is made by the reaction of tin with hydrogen fluoride according to the following equation.  
$$\text{Sn} + 2\text{HF} \rightarrow \text{SnF}_2 + \text{H}_2$$

How many grams of  $\text{SnF}_2$  are produced from the reaction of 30 g of  $\text{HF}$  with  $\text{Sn}$ ? **118 g**
7. In a spacecraft, the carbon dioxide exhaled by astronauts can be removed by its reaction with lithium hydroxide,  $\text{LiOH}$ , according to the following chemical equation.  
$$\text{CO}_2(\text{g}) + 2\text{LiOH}(\text{s}) \rightarrow \text{Li}_2\text{CO}_3(\text{s}) + \text{H}_2\text{O}(\text{l})$$

If 448 L of  $\text{CO}_2$  is exhaled, the average amount exhaled by a person each day, how much (in grams)  $\text{Li}_2\text{CO}_3(\text{s})$  is produced? **1480g  $\text{Li}_2\text{CO}_3$**
8. When 9.8g of aluminum oxide decomposes, how many grams of aluminum metal are produced? **5.2 g Al**
9. How many grams of iodine are produced when 0.72 mol of fluorine react with potassium iodide?
10. How many grams of sodium are required to react with water to produce 5.0g of sodium hydroxide? ( $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$ )
11. How many grams of  $\text{H}_2$  are produced in the reaction in #10?
12. What volume would be occupied by the resulting hydrogen at STP?
13. 20.0g of  $\text{K}$  react with  $\text{H}_2\text{O}$  to produce  $\text{KOH}$  and  $\text{H}_2$  gas. How many liters would the hydrogen gas occupy at STP?
14. If 30.2g of  $\text{Al}$  react with  $\text{HCl}$  to produce  $\text{AlCl}_3$  and  $\text{H}_2$  gas, how many liters of  $\text{H}_2$  are produced at STP?
15. How many liters of  $\text{HCl}$  are produced by the reaction of 5.7L of  $\text{H}_2$  with  $\text{Cl}_2$ ?

# Mixed Stoichiometry Review Worksheet Answers

**E Durkheim**



## **Mixed Stoichiometry Review Worksheet Answers:**

## Whispering the Strategies of Language: An Psychological Journey through **Mixed Stoichiometry Review Worksheet Answers**

In a digitally-driven earth where monitors reign great and quick conversation drowns out the subtleties of language, the profound techniques and psychological subtleties hidden within phrases often go unheard. Yet, located within the pages of **Mixed Stoichiometry Review Worksheet Answers** a charming fictional value sporting with natural emotions, lies an extraordinary quest waiting to be undertaken. Composed by an experienced wordsmith, this marvelous opus invites readers on an introspective journey, delicately unraveling the veiled truths and profound impact resonating within the very fabric of each word. Within the emotional depths of this emotional review, we can embark upon a genuine exploration of the book is key subjects, dissect their interesting writing fashion, and fail to the powerful resonance it evokes deep within the recesses of readers hearts.

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