Gas Laws Worksheet

atm = 760.0 mm Hg = 101.3 kPa= 760.0 torr

Boyle's Law Problems:

PV. TRV

 If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(749-4)(22.51) = (725-4) V_2$$

 $V_2 = \frac{(749-4)(22.51)}{(725-4)}$

V2 - 23.2L

A gas with a volume of 4.0L at a pressure of 205kPa is allowed to expand to a volume of 12.0L.
 What is the pressure in the container if the temperature remains constant?

 What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

 A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

Mixed Gas Law Calculation Work For Problems

Arturo Cuomo

Mixed Gas Law Calculation Work For Problems:

CliffsNotes Chemistry Practice Pack Charles Henrickson, 2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level US Navy diving manual ,1981

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