

1st year Chemistry Guess paper

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Important Short Questions

Chapter# 1

1. How N_2 and CO have same number of electrons, protons and neutrons?
2. No individual neon atom in the sample of the element has a mass of 20.18 amu. Justify.
3. Atomic masses are in fractions. Justify.
4. How limiting reactant controls the amount of product formed?
5. 32 g of O_2 , 28 g of N_2 and 16 g of CH_4 have different masses but same volume. Why?
6. Why actual yield is always lesser than theoretical yield?
7. Law of conservation of mass has to be obeyed during stoichiometric calculations. Explain.
8. One mole of H_2SO_4 should completely react with two moles of NaOH. How does Avogadro's Number helps us too explain it.
9. Calculate the mass of 5.136 moles Ag_2CO_3 .
10. What are molecular ions. How they are formed. Give their applications.
11. How the efficiency of a chemical reaction is determined?

Chapter # 2

12. Why NaCl cannot be purified by crystallization?
13. How crystals are dried? Vacuum Desiccator is safe and reliable method of drying the crystals.
14. How coloured impurities are removed from crystals during crystallization.
15. Give uses of chromatography.
16. Differentiate between adsorption and partition chromatography.
17. Define 'Sublimation' with an example.
18. Why concentrated HCl and $KMnO_4$ cannot be filtered by Gooch crucible?
19. Repeated extractions using small portions of solvent are more efficient than using a single extraction. Why?

Chapter # 3

20. Gases behave ideally at low pressure and high temperature. Explain.
21. Calculate value of 'R' SI units.
22. Water vapours do not behave at 273K. Why?

Intermediate 1st Year Chemistry Guess Paper 2013

John T. Lett, Howard Adler, Max Zelle



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This book is an invaluable resource for teachers across all content areas masters and doctoral research method classes and future Teachers of the Year

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